

Honors Chemistry Rate Law Homework

- 1) For $2A + B \rightarrow C$, we've determined the following experimental data:

Trial	[A] (M)	[B] (M)	Initial rate (M/s)
1	0.0100	0.0100	1.62×10^{-5}
2	0.0200	0.0100	3.24×10^{-5}
3	0.0100	0.0200	6.48×10^{-5}

Using this information, determine the rate law for this reaction, find the reaction order for A and B, and find the overall reaction order.

- 2) In the reaction $A + B \rightarrow C$, we found the following rate data:

Trial	[A] (M)	[B] (M)	Initial rate (M/s)
1	0.026	0.015	2.80×10^{-3}
2	0.026	0.030	5.60×10^{-3}
3	0.052	0.015	11.2×10^{-3}

Using this information, find the overall rate law for this reaction, find the order of the reaction for each reactant, and the overall reaction order.

- 3) Using the information from the rate equation you found in problem 2, determine the rate constant k for this reaction.

Honors Chemistry Rate Law Homework

G Thomas



Honors Chemistry Rate Law Homework:

Enjoying the Track of Term: An Mental Symphony within **Honors Chemistry Rate Law Homework**

In a global eaten by screens and the ceaseless chatter of fast transmission, the melodic beauty and mental symphony developed by the prepared term often diminish in to the background, eclipsed by the persistent sound and disruptions that permeate our lives. But, nestled within the pages of **Honors Chemistry Rate Law Homework** a marvelous fictional value full of organic emotions, lies an immersive symphony waiting to be embraced. Constructed by an outstanding composer of language, that fascinating masterpiece conducts viewers on a mental journey, skillfully unraveling the concealed melodies and profound affect resonating within each cautiously crafted phrase. Within the depths of the touching review, we shall investigate the book is main harmonies, analyze their enthralling writing design, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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