

*Honors Chemistry Midterm: Review Packet***PART 1**

1. Write in the number of significant figures for each of the following measurements:

3 3.00 mL 3 0.00290 g 4 50.00 m 2 0.070 kg 1 400 L

2. Using the following masses: 0.0562 kg, 124.213 g and 1635 mg.
- ← convert to grams*
-
- The answer should be reported as:
-
- g

56.2g 124.213g 1.635g

3. a) A 5.75 mL sample of mercury has a measured mass of 77.05 g. The density is
- 13.49/mL

$$D = \frac{m}{V} = \frac{77.05g}{5.75mL} =$$

- b) Mercury's accepted density is 13.53 g/mL. The % error in your measurement is:
- .9617%

$$\frac{|\text{actual} - \text{exp}|}{\text{actual}} \times 100 \quad \frac{|13.53g/mL - 13.49g/mL|}{13.53g/mL} \times 100$$

4. In the diagram below:



Is this person accurate? no - not in center

Is this person precise? yes - close together

5. Identify the following as a chemical change (CC), physical change (PC), chemical property (CP), or physical property (PP):

a) burning wood CC b) mass of wood PP
c) flammability of wood CP d) chopping up wood PC

Honors Chemistry Blue Packet Answers

Robert R. Redfield



Honors Chemistry Blue Packet Answers:

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Table of Contents Honors Chemistry Blue Packet Answers

1. Understanding the eBook Honors Chemistry Blue Packet Answers
 - The Rise of Digital Reading Honors Chemistry Blue Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Honors Chemistry Blue Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Honors Chemistry Blue Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Honors Chemistry Blue Packet Answers
 - Personalized Recommendations
 - Honors Chemistry Blue Packet Answers User Reviews and Ratings

- Honors Chemistry Blue Packet Answers and Bestseller Lists
- 5. Accessing Honors Chemistry Blue Packet Answers Free and Paid eBooks
 - Honors Chemistry Blue Packet Answers Public Domain eBooks
 - Honors Chemistry Blue Packet Answers eBook Subscription Services
 - Honors Chemistry Blue Packet Answers Budget-Friendly Options
- 6. Navigating Honors Chemistry Blue Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Honors Chemistry Blue Packet Answers Compatibility with Devices
 - Honors Chemistry Blue Packet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Honors Chemistry Blue Packet Answers
 - Highlighting and Note-Taking Honors Chemistry Blue Packet Answers
 - Interactive Elements Honors Chemistry Blue Packet Answers
- 8. Staying Engaged with Honors Chemistry Blue Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Honors Chemistry Blue Packet Answers
- 9. Balancing eBooks and Physical Books Honors Chemistry Blue Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Honors Chemistry Blue Packet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Honors Chemistry Blue Packet Answers
 - Setting Reading Goals Honors Chemistry Blue Packet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Honors Chemistry Blue Packet Answers
 - Fact-Checking eBook Content of Honors Chemistry Blue Packet Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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