

Study Guide for Honors Chemistry Final - June 2015

You will be given all needed equations, values of constants and other pertinent information, including the following: periodic table, $c = \lambda\nu$, $E = h\nu$, the value for $h = 6.626 \times 10^{-34}$ J·s, the conversion between meters and nanometers ($1 \text{ m} = 1 \times 10^9 \text{ nm}$), electronegativity values, activity series of metals, gas law equations, dilution equation, molarity, molality. It is your responsibility to recognize the equations and understand how to use them. You will also be given a 3" x 5" notecard that, must be handwritten by yourself, you may use to organize your thoughts for the exam.

Do the sample problems and review all the vocabulary words in the sections indicated. Go to the following links to test your knowledge on the Chapters (remember we did not do all of the sections in all of the following chapters):

http://chemistrymc.com/chapter_test and http://chemistrymc.com/standardized_test

Electrons in Atoms – Chapter 5

Objectives:

- Describe the relationship between wavelength and frequency.
- Be able to use the equation $c = \lambda\nu$ where $c = 3.00 \times 10^8$ m/s
- Describe the relationship between frequency and energy of a wave.
- Be able to use the equation $E = h\nu$ where $h = \text{Planck's constant} = 6.626 \times 10^{-34}$ J·s
- Describe the contributions of Bohr, Schrödinger, Planck, Einstein, & deBroglie to determining the location of electrons in the atom and their properties
- Explain the difference between a continuous spectrum and a line spectrum and give examples of each.
- Be able to arrange the colors of the visible spectrum from lowest energy to highest energy.
- Define a quantum of energy and explain how it is related to an energy change of matter.
- Describe how the quantum mechanical model describes what an atom is and where the subatomic particles are located.
- Explain the relationships among an atom's energy level, sublevels, and atomic orbitals.
- Apply the Pauli Exclusion Principle, the Aufbau Principle, and Hund's Rule to write electron configurations using orbital diagrams and electron configuration notation.
- Define valence electrons and draw electron-dot structures representing an atom's valence electrons.
- Be able to read and write quantum numbers given an electron configuration or an element's symbol
- Labs – Thickness of Zinc, Spectroscopy and Flame Tests, Electron Arrangements

Vocabulary:

- | | |
|---|--|
| • Wavelength, frequency, amplitude and speed of waves | • Principal energy level |
| • Electromagnetic spectrum | • Orbitals (s, p, d, and f) |
| • Quantum | • Electron configuration |
| • Photoelectric effect | • Valence electron |
| • Atomic emission spectrum | • Electron-dot structures (Lewis dot diagrams) |
| • Ground state vs. excited state | • Quantum numbers |
| • Heisenberg uncertainty principle | • $n, \ell, m\ell, m_s$ |
| • Quantum mechanical model of the atom | |
| • Atomic orbital | |
| • Principal quantum number | |

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Siddappa N.Byrareddy



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Table of Contents Honors Chemistry Final Study Guide

1. Understanding the eBook Honors Chemistry Final Study Guide
 - The Rise of Digital Reading Honors Chemistry Final Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Honors Chemistry Final Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Honors Chemistry Final Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Honors Chemistry Final Study Guide
 - Personalized Recommendations
 - Honors Chemistry Final Study Guide User Reviews and Ratings
 - Honors Chemistry Final Study Guide and Bestseller Lists

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 - Honors Chemistry Final Study Guide Budget-Friendly Options
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 - Honors Chemistry Final Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Honors Chemistry Final Study Guide
 - Highlighting and Note-Taking Honors Chemistry Final Study Guide
 - Interactive Elements Honors Chemistry Final Study Guide
8. Staying Engaged with Honors Chemistry Final Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Honors Chemistry Final Study Guide
9. Balancing eBooks and Physical Books Honors Chemistry Final Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Honors Chemistry Final Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Honors Chemistry Final Study Guide
 - Setting Reading Goals Honors Chemistry Final Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Honors Chemistry Final Study Guide
 - Fact-Checking eBook Content of Honors Chemistry Final Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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