

Equilibria Involving A Weak Base

You have 0.010 M NH_3 . Calc. the pH.



$$K_b = 1.8 \times 10^{-5}$$

Step 1. Define equilibrium concs. in ICE table

	$[\text{NH}_3]$	$[\text{NH}_4^+]$	$[\text{OH}^-]$
initial	0.010	0	0
change	-x	+x	+x
equilib	0.010 - x	x	x

Honors Chemistry Equilibrium

Liyong Dong



Honors Chemistry Equilibrium:

Embracing the Beat of Expression: An Mental Symphony within **Honors Chemistry Equilibrium**

In a global used by displays and the ceaseless chatter of instant connection, the melodic splendor and emotional symphony created by the published term usually diminish in to the backdrop, eclipsed by the constant noise and interruptions that permeate our lives. However, set within the pages of **Honors Chemistry Equilibrium** an enchanting fictional prize overflowing with fresh thoughts, lies an immersive symphony waiting to be embraced. Constructed by an outstanding musician of language, that interesting masterpiece conducts visitors on a mental trip, well unraveling the hidden songs and profound impact resonating within each cautiously crafted phrase. Within the depths of this touching analysis, we shall examine the book is main harmonies, analyze their enthralling publishing type, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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