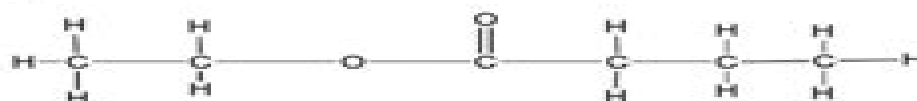


1.1	B	✓	✓
1.2	B	✓	✓
1.3	A	✓	✓
1.4	B	✓	✓
1.5	A	✓	✓
1.6	A	✓	✓
1.7	D	✓	✓
1.8	D	✓	✓
1.9	C	✓	✓
1.10	D	✓	✓

(2 × 10 = 20)

2.1



✓

✓

✓

(3)

2.2 They are volatile and evaporate ✓ quickly. (1)

2.3 ethanol ✓ and butanoic acid ✓ (2)

2.4 hydrolysis/elimination/esterification/condensation (any one) ✓ (1)

2.5 heat ✓ and sulphuric acid (H<sub>2</sub>SO<sub>4</sub>) ✓ (2)

3.1 Methane is a non-polar substance ✓ and is therefore insoluble ✓ in water. Methanol is polar and contains an -OH group ✓ which allows it to make hydrogen bonds ✓. (4)

3.2 Methane and Methanal ✓ ✓ (2)

3.3 As the mass of the carboxylic acid increases, the boiling point of the carboxylic acids also increases ✓. This is because larger molecules have more intermolecular forces ✓ than smaller molecules. (2)

4.1.1 Hydration ✓ (1)

4.1.2 No ✓ because it needs a catalyst and heat ✓ to take place. (2)

4.1.3 propan ✓ -1-ol ✓ (2)

4.1.4 primary ✓ (1)

4.1.5 No ✓, also propan-2-ol ✓ (2)

4.2.1 hydrohalogenation ✓ (1)

4.2.2 no water present ✓ (1)

# Grade Memo September 2014

**Jessica J Manson**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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