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Imstructionse

- * Periodic Table is provided.
- This paper consists of 09 pages.
- · Answer all the questions.
- Use of calculators is not allowed.
- Write your Index Number in the space provided in the answer sheet.
- Follow the instructions given on the back of the answer sheet carefully.
- * In each of the questions 1 to 50, pick one of the alternatives from (1), (2), (3), (4), (5) which is correct or most appropriate and mark your response on the answer sheet with a cross (x) in accordance with the instructions given on the back of the answer sheet.

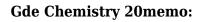
Universal gas constant $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$ Avogadro constant $N_a = 6.022 \times 10^{23} \text{ mol}^{-1}$ Planck's constant $h = 6.626 \times 10^{-34} \text{ J s}$ Velocity of light $c = 3 \times 10^8 \text{ m s}^{-1}$

- I. Which of the following electronic transitions belongs to the visible region of the line spectrum of atomic hydrogen? (n = Principal quantum number)
 - (1) $n=5 \rightarrow n=3$ (2) $n=4 \rightarrow n=2$
 - (3) $n = 1 \rightarrow n = 2$
 - (4) $n=3 \rightarrow n=1$
 - (5) $n=2 \rightarrow n=1$
- 2. Select the incorrect statement.
 - (1) Pauli exclusion principle excludes the possibility of the presence of a third electron in an orbital.
 - (2) In a potassium atom, the number of electrons having quantum numbers n (principal quantum number) = 3 and m, (magnetic quantum number) = 0 is 4.
 - (3) The effective nuclear charge felt by a valence electron in nitrogen (N), is greater than the effective nuclear charge felt by a valence electron in carbon (C).
 - (4) Among the ions Na*, Mg^{2*}, K* and Ca^{2*}, the two ions closest to each other in size are K* and Mg^{2*}.
 - (5) Electron gain energy of carbon is negative.
- The increasing order of the second ionization energy (X*(g) → X²*(g) + e) of Be, B and O is,
 - (I) Be < B < O
 - (2) Be < O < B (3) B < O < Be
 - (4) B < Be < O
 - (5) O < Be < B
- 4. The shapes of F₄ClO, FClO, and FClO, are respectively,
 - (1) tetrahedral, trigonal planar and see-saw.
 - (2) square planar, trigonal planar and tetrahedral.
 - (3) see-saw, trigonal pyramidal and square planar.
 - (4) tetrahedral, trigonal pyramidal and see-saw.
 - (5) see-saw, trigonal pyramidal and tetrahedral.

[See page two

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