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 இயற்பியல் I  
 Chemistry I

02 E I

පැය තුනයි  
 இரண்டு மணித்தியாலம்  
 Two hours

### Instructions:

- Periodic Table is provided.
- This paper consists of 09 pages.
- Answer all the questions.
- Use of calculators is not allowed.
- Write your Index Number in the space provided in the answer sheet.
- Follow the instructions given on the back of the answer sheet carefully.
- In each of the questions 1 to 50, pick one of the alternatives from (1), (2), (3), (4), (5) which is correct or most appropriate and mark your response on the answer sheet with a cross (x) in accordance with the instructions given on the back of the answer sheet.

Universal gas constant  $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$

Planck's constant  $h = 6.626 \times 10^{-34} \text{ J s}$

Avogadro constant  $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$

Velocity of light  $c = 3 \times 10^8 \text{ m s}^{-1}$

- Which of the following electronic transitions belongs to the visible region of the line spectrum of atomic hydrogen? ( $n$  = Principal quantum number)
  - $n = 5 \rightarrow n = 3$
  - $n = 4 \rightarrow n = 2$
  - $n = 1 \rightarrow n = 2$
  - $n = 3 \rightarrow n = 1$
  - $n = 2 \rightarrow n = 1$
- Select the incorrect statement.
  - Pauli exclusion principle excludes the possibility of the presence of a third electron in an orbital.
  - In a potassium atom, the number of electrons having quantum numbers  $n$  (principal quantum number) = 3 and  $m_l$  (magnetic quantum number) = 0 is 4.
  - The effective nuclear charge felt by a valence electron in nitrogen (N), is greater than the effective nuclear charge felt by a valence electron in carbon (C).
  - Among the ions  $\text{Na}^+$ ,  $\text{Mg}^{2+}$ ,  $\text{K}^+$  and  $\text{Ca}^{2+}$ , the two ions closest to each other in size are  $\text{K}^+$  and  $\text{Mg}^{2+}$ .
  - Electron gain energy of carbon is negative.
- The increasing order of the second ionization energy ( $\text{X}^+(\text{g}) \rightarrow \text{X}^{2+}(\text{g}) + e$ ) of Be, B and O is,
  - $\text{Be} < \text{B} < \text{O}$
  - $\text{Be} < \text{O} < \text{B}$
  - $\text{B} < \text{O} < \text{Be}$
  - $\text{B} < \text{Be} < \text{O}$
  - $\text{O} < \text{Be} < \text{B}$
- The shapes of  $\text{F}_3\text{ClO}$ ,  $\text{FCIO}_2$  and  $\text{FCIO}_3$  are respectively,
  - tetrahedral, trigonal planar and see-saw.
  - square planar, trigonal planar and tetrahedral.
  - see-saw, trigonal pyramidal and square planar.
  - tetrahedral, trigonal pyramidal and see-saw.
  - see-saw, trigonal pyramidal and tetrahedral.

[See page two]

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