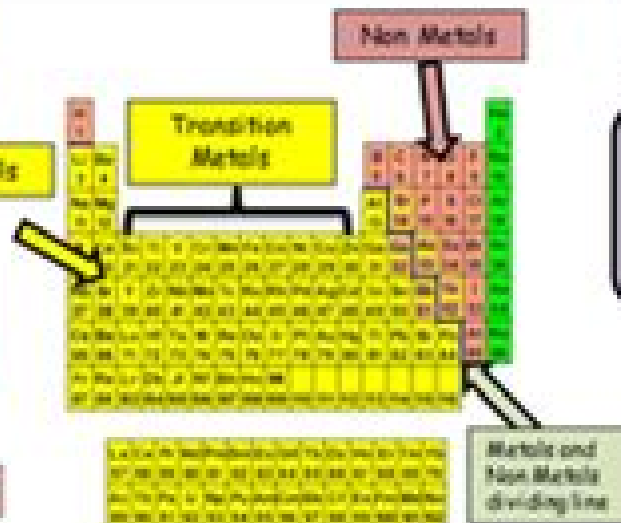


Year 11 C2 - Discovering Chemistry Topic 1



Particle	Relative Mass	Relative Charge
Proton	1	+
Neutron	1	0
Electron	Very small	-



I came up with the first periodic table. I arranged the elements in order of increasing atomic mass

Dmitri Mendeleev



I put elements with similar properties into horizontal rows. I also left gaps incase any new elements were discovered

He

4 ← **MASS NUMBER** = number of protons + number of neutrons

2 ← **PROTON NUMBER** = number of protons (atomic)

CHARGE

Isotopes

Higher

Example is an atom with a different number of neutrons

Notice that the mass number is different. How many neutrons does each isotope have?

16 17 18

Each isotope has 8 protons - if it didn't then it just wouldn't be oxygen any more

14
N
7



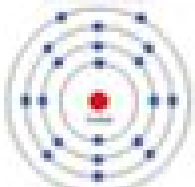
Electron structure
+ 2, 5

24
Mg
12

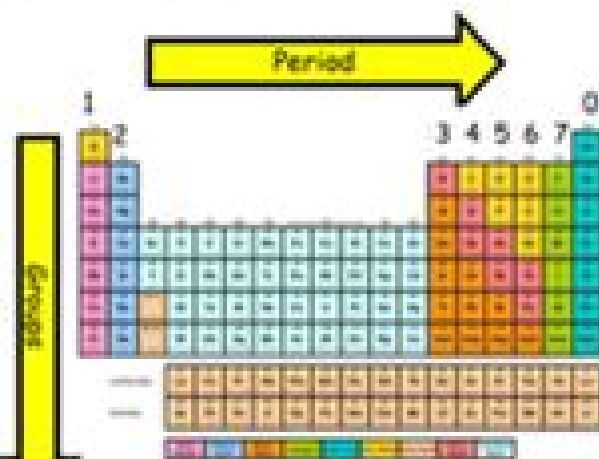


Electron structure
+ 2, 8, 2

40
Ca
20



Electron structure
+ 2, 8, 8, 2

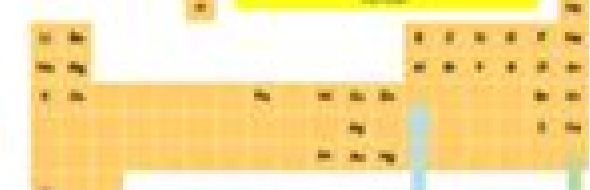


How can an atom have a decimal for its mass?

This is because out of every four naturally occurring chlorine atoms, 3 have a mass of 35 and 1 has a mass of 37 so the average atomic mass is:

$$35 \times 3 + 1 \times 37 = 106$$

Elements in the same group have the same number of electrons in the outer shell (this corresponds to their group number)



All of group 1 metals have 1 electron in their outer shell

These elements have 3 electrons in their outer shell

These elements have 8 electrons in their outer shell

(Most important) All of the elements in the same group have similar **PROPERTIES**. This is how I thought of the periodic table in the first place. This is called **PERIODICITY**.

May 2013 Chemistry C2

M Lipman



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