

C2 - THE PERIODIC TABLE 2.6 The transition metals

GCSE Chemistry Aiming for Grade 4

Where are the transition metals on the periodic table?
In the centre block

Give the names of five transition metals

Gold, Silver, Aluminium, Iron, Copper

List the physical properties of the transition metals

1. Good conductors of heat
2. Good conductors of electricity
3. High densities
4. High melting points
5. Hard and strong

Why is mercury not typical as a transition metal?

It is a liquid at room temperature
whereas all the others are solid with
high melting points.



Assessment
Grade 4

GCSE Chemistry Aiming for Grade 6

Compared to Group 1 metals how do the following
properties compare?

Density

**Much higher than group 1, none of them will float on
water like lithium, sodium and potassium.**

Melting point

Much higher melting points except for mercury

Boiling point

**Boiling points are extremely high compared to the
alkali metals**

Hardness

**Alkali metals can be cut with a knife and are soft.
Transition metals are very hard.**

Strength

**Transition metals are very strong and can be used for
buildings and machines due to their strength**

Give the names of the following transition metal
compounds.

Fe_2O_3

Iron Oxide

Cu_2O

Copper Oxide

NiCO_3

Nickel Carbonate

AgNO_3

Silver Nitrate

Assessment
Grade 6

GCSE Chemistry Aiming for Grade 8

Why is copper used for wires?
It is a very good conductor -
other metals. It is also very
drawn into wires.

Steel is made from Iron and
make girders on bridges?
It is very strong and hard to
weight and high forces.

Titanium is used for aircraft
Titanium has a very high melting
point ideal for very hot engine parts

State the colours of the compounds
of transition metals.

Copper sulphate **blue** due to

Nickel Carbonate **pale green**

Chromium oxide **dark green**

Manganese Chloride **pale pink**

Compared with Group 1 elements
metals have different properties
Transition metals can form a
variety of ions as their reactivity is different.
eg Iron can form Fe^{2+} and Fe^{3+}
Different reactivity.

Assessment
Grade 8

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