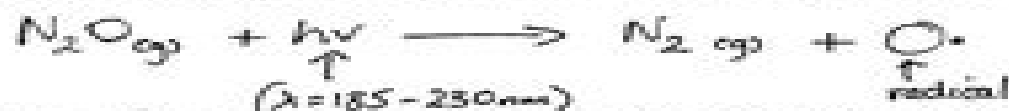


F332 advance notice (June 2013) "predicted Q's"

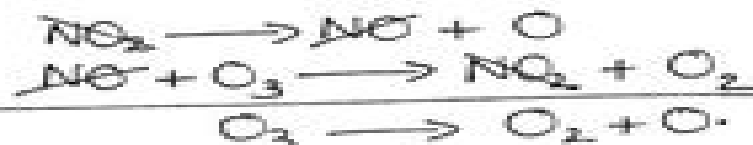
- How does photolysis play a part in the destruction of N_2O in the stratosphere?



[90% of N_2O loss in Stratosphere is this!]

* THIS IS "PHOTODISSOCIATION". $O\cdot$ helps to form tropospheric O_3 .

- Use a series of equations to show how NO acts in the catalytic destruction of ozone?



(overall)

This can be a bad thing at tropospheric level. NO can react with C_xH_y , preventing NO breaking down O_3 . Build up of O_3 in Trop. leads to photochemical smog!!

- What is meant by the term photolysis?

"photo" = LIGHT

"-lysis" = SPLITTING

Light energy drives chemical reactions

- Convert 100 parts per billion to a percentage, for the level of NO_x in urban areas

1 part per million $\longrightarrow$ $\frac{1}{10,000}$ (divide by 10,000)

100 ppb = 0.1 ppm (1,000 ppb = 1 ppm)

$$\frac{0.1 \text{ ppm}}{10,000} = \underline{\underline{1 \times 10^{-5} \%}} \text{ (or } 0.00001 \%)$$

- Why is the $N_2O_{(atmos)} \rightleftharpoons N_2O_{(ocean)}$ not a true dynamic equilibrium?

ATMOSPHERE + OCEANS can't be regarded as ~~as~~ a "SEALED system".
(gases "lost").

- Describe how photochemical smog can be harmful to:
(a) the environment?

- crop damage

- compounds with $C=C$ bonds are particularly vulnerable
(rubbers, paints ...)

June 2013 F332 Paper

Gerhard L. Weinberg



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