

Section A

For each question there are four possible answers, A, B, C and D. Choose the one you consider to be correct.

- 1 Which equation correctly represents the enthalpy change of atomisation of iodine?



- 2 What is the second electron affinity of oxygen?
Use the data below to determine your answer.

	$\Delta H \text{ kJ/mol}$
Enthalpy of formation of calcium oxide	-636
Lattice enthalpy of calcium oxide	-3401
Second ionisation energy of calcium	+1100
First ionisation energy of calcium	+590
Enthalpy of atomisation of calcium	+177
Enthalpy of atomisation of oxygen	+249
First electron affinity of oxygen	-141

A $+790 \text{ kJ/mol}$

B -790 kJ/mol

C $+1\,039 \text{ kJ/mol}$

D $-1\,039 \text{ kJ/mol}$

- 3 Why does the enthalpy change of neutralisation of oleic acid, $\text{CH}_3(\text{CH}_2)_7\text{CH}=\text{CH}(\text{CH}_2)_7\text{CO}_2\text{H}$, with aqueous ammonia differ from the theoretical value of $-57.2 \text{ kJ/mol}^{-1}$?

A Oleic acid and ammonia do not react in the ratio 1:1.

B Oleic acid is unsaturated therefore its enthalpy of neutralisation is more exothermic.

C Energy is required to ionise oleic acid and ammonia.

D Ammonia is completely hydrolysed.

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Yan Bai

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Table of Contents June 2013 Chemistry Paper 3

1. Understanding the eBook June 2013 Chemistry Paper 3
 - The Rise of Digital Reading June 2013 Chemistry Paper 3
 - Advantages of eBooks Over Traditional Books
2. Identifying June 2013 Chemistry Paper 3
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an June 2013 Chemistry Paper 3
 - User-Friendly Interface
4. Exploring eBook Recommendations from June 2013 Chemistry Paper 3
 - Personalized Recommendations
 - June 2013 Chemistry Paper 3 User Reviews and Ratings
 - June 2013 Chemistry Paper 3 and Bestseller Lists
5. Accessing June 2013 Chemistry Paper 3 Free and Paid eBooks

- June 2013 Chemistry Paper 3 Public Domain eBooks
 - June 2013 Chemistry Paper 3 eBook Subscription Services
 - June 2013 Chemistry Paper 3 Budget-Friendly Options
6. Navigating June 2013 Chemistry Paper 3 eBook Formats
 - ePub, PDF, MOBI, and More
 - June 2013 Chemistry Paper 3 Compatibility with Devices
 - June 2013 Chemistry Paper 3 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of June 2013 Chemistry Paper 3
 - Highlighting and Note-Taking June 2013 Chemistry Paper 3
 - Interactive Elements June 2013 Chemistry Paper 3
 8. Staying Engaged with June 2013 Chemistry Paper 3
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers June 2013 Chemistry Paper 3
 9. Balancing eBooks and Physical Books June 2013 Chemistry Paper 3
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection June 2013 Chemistry Paper 3
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine June 2013 Chemistry Paper 3
 - Setting Reading Goals June 2013 Chemistry Paper 3
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of June 2013 Chemistry Paper 3
 - Fact-Checking eBook Content of June 2013 Chemistry Paper 3
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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