

Question 1

(a) Using only the periodic table, arrange each set of atoms in order of increasing radius:

(1) Ba, Ca, Na



(2) In, Sn, As



(3) Cs, K, Al



(b) Which neutral atom is isoelectronic with each of the following ions?

(1) H^+

None

(The H^+ ion has 0 electrons. It is impossible for a neutral atom to have 0 electrons.)

(2) Zr^{4+}

Kr

(3) Mn^{7+}

Ar

(4) I^-

Xe

(c) Explain the following variations in atomic or ionic radii:

(1) $\text{I}^- > \text{I} > \text{I}^+$

As the atomic number stays constant and the number of electrons increases, the electron-electron repulsions increase. This leads to the electrons spreading out, and the anion becomes larger. The reverse is also true, so the cation is smaller.

(2) $\text{Ca}^{2+} > \text{Mg}^{2+} > \text{Be}^{2+}$

For cations with the same charge, ionic radii increase going down a column because of the increase in principal quantum number, which means the outer electrons are farther away from the nucleus on average. This leads to larger ion size.

Ionic Radius Answer Key

M Carnoy

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