

Name Answer Key Hour \_\_\_\_\_

## Chapter 4 Review (Honors Chemistry)

Use the given number of neutrons and electrons to fill in the chart and determine whether the atom is...

Common: all information is exactly like the information on the periodic table.

Ion: the electron number differs from the proton number, but all other information is the same.

Isotope: the mass of the atom is different than the most common version.

|     | ELEMENT   | ATOMIC NUMBER | MASS | PROTON NUMBER | NEUTRON NUMBER | ELECTRON NUMBER | ISOTOPE, ION OR COMMON |
|-----|-----------|---------------|------|---------------|----------------|-----------------|------------------------|
| 1.  | Sc        | 21            | 45   | 21            | 24             | 18              | Cation, $Sc^{3+}$      |
| 2.  | K         | 19            | 42   | 19            | 23             | 19              | Isotope                |
| 3.  | Cobalt    | 27            | 59   | 27            | 32             | 29              | Anion, $Co^{2-}$       |
| 4.  | S         | 16            | 32   | 16            | 16             | 18              | Anion, $S^{2-}$        |
| 5.  | Argon     | 18            | 40   | 18            | 22             | 18              | Common                 |
| 6.  | Uranium   | 92            | 238  | 92            | 146            | 92              | Common                 |
| 7.  | Ca        | 20            | 43   | 20            | 23             | 20              | Isotope                |
| 8.  | P         | 15            | 31   | 15            | 16             | 18              | Anion, $P^{3-}$        |
| 9.  | Pb        | 82            | 225  | 82            | 143            | 82              | Isotope                |
| 10. | Zinc      | 30            | 65   | 30            | 35             | 30              | Common                 |
| 11. | Aluminum  | 13            | 30   | 13            | 17             | 13              | Isotope                |
| 12. | Potassium | 19            | 39   | 19            | 20             | 18              | Cation, $K^{+}$        |
| 13. | Ce        | 58            | 142  | 58            | 84             | 58              | Isotope                |
| 14. | Sodium    | 11            | 23   | 11            | 12             | 10              | Cation, $Na^{+}$       |
| 15. | Chlorine  | 17            | 35   | 17            | 18             | 18              | Anion, $Cl^{-}$        |
| 16. | Radon     | 86            | 222  | 86            | 136            | 86              | Common                 |
| 17. | Mg        | 12            | 22   | 12            | 10             | 12              | Isotope                |
| 18. | Boron     | 5             | 11   | 5             | 6              | 5               | Common                 |
| 19. | Nickel    | 28            | 59   | 28            | 31             | 26              | Cation, $Ni^{2+}$      |
| 20. | Bismuth   | 83            | 209  | 83            | 126            | 83              | Common                 |
| 21. | Iridium   | 77            | 196  | 77            | 119            | 77              | Isotope                |
| 22. | Hg        | 80            | 201  | 80            | 121            | 78              | Cation, $Hg^{2+}$      |
| 23. | Titanium  | 22            | 50   | 22            | 28             | 22              | ISOTOPE                |

# Honors Chemistry Review Study Guide Key

**Jin-Ying Zhang**



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