

100

27. It is necessary to know the formula of a compound in order to calculate its percentage composition.

101

28. If the percent by mass of carbon in methane, CH_4 , is 75%, then 100 grams of methane contains 75 g of carbon and 25 g of hydrogen.

102

29. The formula for methane, CH_4 , is both a molecular and an empirical formula.

103

30. The empirical formula of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, is CH_2O .

C. Problems

Solve the following problems on the space provided. Show your work.

31. How many moles of Fe in 5.1 g of 10^{23} moles of Fe?

32. How many representative particles are in 2.18 g 10^{23} and MgCl_2 ?

33. What is the gram molecular mass of methane, CH_4 ?

34. Find the mass of each of the following:

a. 2.43 g 10^{23} and H_2O_2

b. 2.43 g and H_2O_2

c. 4.30 g and H_2O_2

d. 4.30 g and H_2O_2

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