

HALF-LIFE OF RADIOACTIVE ISOTOPES

Name _____

1. How much of a 100.0 g sample of ^{198}Au is left after 8.10 days if its half-life is 2.70 days?

$$\frac{8.1}{2.7} = 3$$

$$\frac{100}{2} = 50$$

$$\frac{50}{2} = 25$$

$$\frac{25}{2} = 12.5$$

12.5 g

2. A 50.0 g sample of ^{14}N decays to 12.5 g in 14.4 seconds. What is its half-life?



7.7 sec

3. The half-life of ^{40}K is 12.4 hours. How much of a 750 g sample is left after 62.0 hours?

$$\frac{62}{12.4} = 5$$

$$\left(\frac{750}{2}\right) = 375$$

$$\left(\frac{375}{2}\right) = 187.5$$

$$\frac{187.5}{2} = 93.75$$

93.75 g

4. What is the half-life of ^{99}Tc if a 500 g sample decays to 62.5 g in 639,000 years?

$$\frac{500}{8} = 62.5$$

$$\left(\frac{1}{2}\right)^n = \frac{1}{8} \Rightarrow n = 3$$

$$\frac{639,000}{3} = 213,000$$

213,000 yr

5. The half-life of ^{232}Th is 1.4×10^{10} years. If there are 25.0 g of the sample left after 2.8×10^{10} years, how many grams were in the original sample?

$$\frac{2.8 \times 10^{10}}{1.4 \times 10^{10}} = 2 \Rightarrow 2^2 = 4$$

$$25.0 \cdot 4 = 100.0$$

100.0 g

6. There are 5.0 g of ^{131}I left after 40.35 days. How many grams were in the original sample if its half-life is 8.07 days?

$$\frac{40.35}{8.07} = 5$$

$$2^5 = 32$$

$$32 \cdot 5 = 160$$

160 g

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