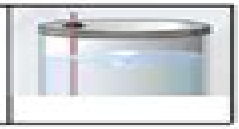
Activity C: Calculating specific heat	Get the Gizmo ready: • Click **Reset**.	

Introduction: The specific heat capacity of a substance is the amount of energy needed to change the temperature of that substance by 1 °C. Specific heat capacity can be calculated using the following equation:

$$q = mc\Delta T$$

In the equation q represents the amount of heat energy gained or lost (in **joules**), m is the mass of the substance (in grams), c is the specific heat capacity of the substance (in J/g °C), and ΔT is the temperature change of the substance (in °C).

Goal: Calculate the specific heat capacities of copper, granite, lead, and ice.

1. Solve: When you mix two substances, the heat gained by one substance is equal to the heat lost by the other substance. Suppose you place 125 g of aluminum in a calorimeter with 1,000 g of water. The water changes temperature by 2 °C and the aluminum changes temperature by −74.95 °C.

   A. Water has a known specific heat capacity of 4.184 J/g °C. Use the specific heat equation to find out how much heat energy the water gained (q).

   $\Delta TQ = 1000 \times 4.18 \times 2Q = 8360\ J8360\ J; 8.4kJ$

   B. Assume that the heat energy gained by the water is equal to the heat energy lost by the aluminum. Use the specific heat equation to solve for the specific heat of aluminum. (Hint: Because heat energy is lost, the value of q is negative.)

   0.893

   Aluminum's accepted specific heat value is 0.900 J/g °C. Use this value to check your work.

2. Calculate: Use the Gizmo to mix 200 g of copper at 100 °C with 1,000 g of water at 20 °C.

   A. What is the final temperature? 21.42

   B. Calculate the temperature change of each substance by subtracting the initial temperature from the final temperature.

   ΔT_{water}: 1.49 ΔT_{copper}: 20

   C. How much heat energy (q) did the water gain? 18

   D. Now solve for the specific heat (c) of copper: it is 20°C

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