Name:	KEY	Date:	Period:

Flame Test Lab Activity Key

Background

The heat from a laboratory burner will cause the ions of some elements to give off light. Electrons will absorb the heat energy from the flame and will "jump" to a higher energy level. When the electrons return to their original energy levels, this absorbed energy is released as light. Different elements absorb and release different amounts of energy and thus produce different kinds of light. This light may be visible or invisible, infrared or ultraviolet radiation. Analysis of this light can be used to identify elements and molecules. In today's lab, we will observe the visible range of light emission.

Objectives

- Perform a flame test to identify the characteristic color of metal ions.
- Identify the metal ion in an unknown solution.
- Calculate the energy of emitted photons.

	Note:	
Materials: (per lab group)	If chloride compounds are not available, metal nitrate	
10 ml dilute solutions of the following: Barium Chloride	compounds may be substituted.	
Calcium Chloride		
Potassium Chloride	Unknowns:	
Sodium Chloride	Number the beakers 1-12.	
Strontium Chloride	Add solutions as follows:	
Unknown Solution		
7 Wood Splints	1. Strontium Chloride	7. Calcium Chloride
1 Bunsen Burner	2. Calcium Chloride	8. Strontium Chloride
Safety Goggles & Aprons	3. Potassium Chloride	9. Barium Chloride
	4. Barium Chloride	Potassium Chloride
	5. Lithium Chloride	11. Sodium Chloride
Procedure	6. Sodium Chloride	12. Lithium Chloride

- Put on your safety goggles and an apron.
- Obtain a wood splint and the barium chloride solution.
- 3. Dip the wood splint into the solution and allow it to soak for 1-2 minutes.
- 4. Light Bunsen Burner, USE CAUTION!
- Remove the splint from the solution and hold the end of the splint into the flame. Do not burn the splint.
- Observe and record the color of the flame.
- Repeat steps 2-6 using the remaining chloride solutions.
- 8. Obtain an unknown solution from your teacher. Record the number of your unknown.
- Soak a wood splint in the unknown solution.
- 10. Hold the splint in the flame and record the color of the flame that is produced.
- Using your data, identify the metal ion in your unknown solution.

Flame Test Lab Metal Ions Answer Key

JG Myers

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