

100) A student finds the ΔG° of hydrolysis of ATP is -30.5 kJ/mol at 37°C in buffered water at pH 7.0. At which condition would hydrolysis of ATP be most spontaneous?

(a) Initial reactants are at their standard states

_____ (b) $[\text{ATP}] = 1 \text{ M}$, $[\text{ADP}] = 1 \text{ M}$, $[\text{P}_i] = 1 \text{ M}$ _____ (30)

(c) Initial is the standard state of reactants of ATP , ADP , and P_i

_____ (d) $[\text{ATP}] = 1 \text{ M}$, $[\text{ADP}] = 1 \text{ M}$, $[\text{P}_i] = 1 \text{ M}$ _____ (30)

101) Calculate the spontaneity of hydrolysis of ATP in a cell at 37°C .

$$\frac{1.2 \times 10^{-3}}{1.2 \times 10^{-3}} = \frac{[\text{ATP}]}{[\text{ADP}][\text{P}_i]} \quad \Delta G^\circ = -30.5 \text{ kJ/mol}$$

_____ (a) $\Delta G^\circ = -30.5 \text{ kJ/mol}$ _____ (b) $\Delta G^\circ = -30.5 \text{ kJ/mol}$ _____ (30)

102) Hydrolysis of ATP is spontaneous under standard conditions with an equilibrium constant $K = 10^5$.

(a) Hydrolysis is spontaneous if ΔG° is negative

_____ (b) $\Delta G^\circ = -30.5 \text{ kJ/mol}$ _____ (30)

103) Phosphorylation of a protein is coupled to the hydrolysis of ATP. The ΔG° for the hydrolysis of ATP is -30.5 kJ/mol .

_____ (a) $\Delta G^\circ = -30.5 \text{ kJ/mol}$ _____ (30)

_____ (b) $\Delta G^\circ = -30.5 \text{ kJ/mol}$ _____ (30)

_____ (c) $\Delta G^\circ = -30.5 \text{ kJ/mol}$ _____ (30)

(Total: 100)



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Thomas Griffiths



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