Consider the isoelectronic species, Na^+ , Mg^{2+} , F^- , and O^{2-} . The correct order of increasing length of their radii is $1. F^- < O^{2-} < Mg^{2+} < Na^+ < 2. Mg^{2+} < Na^+ < F^- < O^{2-} < 3. O^{2-} < F^- < Na^+ < Mg^{2+} < 4. O^{2-} < F^- < Mg^{2+} < Na^+ < Mg^{2+} < 4. O^{2-} < F^- < Mg^{2+} < Na^+ < Mg^{2+} < 4. O^{2-} < F^- < Mg^{2+} < Na^+ < Mg^{2+} < 4. O^{2-} < F^- < Mg^{2+} < Na^+ < Mg^{2+} < Mg^{2+$

Among the following elements, the one that is not an actinoid is: 1. Curium (Z=96) 2. Californium (Z=98)

The order of screening effect of electrons of s, p, d and f orbitals of a given shell of an atom on electrons in its outer shell is:

1. s > p > d > f

2. f > d > p > s 3. p < d < f < s 4. f > p > s > d

Uranium (Z=92).

Terbium (Z=65).

The first ionisation enthalpies of Na, Mg, Al, and Si are in the order of-1. Na < Al < Mg < Si 2. Na > Mg > Al > Si

2: Na > Mg > Al > Si 3. Na < Mg < Al < Si 4. Na > Mg > Al < Si

5 Among halogens, the correct order of amount of energy released in electron gain (electron gain enthalpy) is-

1, F > Cl > Br > I 2, F < Cl < Br < I 3, F < Cl > Br > I 4, F < Cl < Br < I

6 The period number in the long form of the periodic table is equal to:

 Magnetic quantum number of any element of the period.

Atomic number of any element of the period.

Maximum principal quantum number of any element of the period.

 Maximum azimuthal quantum number of any element of the period. The correct order of size of the given species is:

1. I > I' > I'

2. $\Gamma^+ > \Gamma^- > \Gamma$

3. $I > I^+ > I^-$

4. $\Gamma > 1 > 1^+$

8 Electronic configuration of four elements A, B, C

and D are given below:

A. $1s^2 2s^2 2p^6$

B. $1s^2 2s^2 2p^4$

C. $1s^2 2s^2 2p^6 3s^1$

D. $1s^22s^22p^5$

The correct order of increasing tendency to gain electron is:

 $L, A \le C \le B \le D$

2. A < B < C < D

3. $D \le B \le C \le A$

 $4. D \le A \le B \le C$

Match the correct ionization enthalpies and electron

gain enthalpies of the following elements.

	Elements		ΔH_1	ΔH_2	$\Delta_{\rm eg} H$
(i)	Most reactive non-metal	A.	419	3051	-48
(ii)	Most reactive metal	В.	1681	3374	-328
(iii)	Least reactive element	C.	738	1451	50
(iv)	Metal forming binary halide	D.	2372	5251	+48

Codes

	A	В	C	D
1.	iti	i.	in	iši
2.	i	111	101	iw
3.	1	iv	100	10
4.	ïv	i	iiii	111

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Reviewing Mcqs On Inorganic: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Mcqs On Inorganic**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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