



10. Which of the following are endothermic processes?

- A. ice melting
- B. water condensing
- C. freezing
- D. sublimation

11. Heat absorbed per mole at STP, positive temperature change

12. Which of the following is a statement of the first law of thermodynamics?

- A. the negative change in enthalpy corresponds to an exothermic process
- B. energy lost by the system must be gained by the surroundings, neither of which are created
- C. temperature is a measure of average kinetic energy
- D. enthalpy is a state function

13. The specific enthalpy of combustion of glucose is negative. This means that the product is _____

- A. a hydrocarbon with the same enthalpy as the reactant
- B. a hydrocarbon and the temperature of the surroundings will decrease
- C. a hydrocarbon and the surroundings will increase
- D. a hydrocarbon and the temperature of the surroundings will increase
- E. none of these is a true statement

14. The reaction



is _____ and therefore heat is _____ from the system.

- A. endothermic, released
- B. exothermic, absorbed
- C. endothermic, released
- D. exothermic, absorbed
- E. thermodynamic neither released nor absorbed

15. Heat evolved when 1 mole of glucose is oxidized when the 2 and 3 moles of oxygen are added to the reaction
 given below is 1366.8 kJ at 25°C.



- A. 1366.8 kJ
- B. 683.4 kJ
- C. 136.68 kJ
- D. 1366.8 kJ
- E. 136.68 kJ



Mcps Honors Chemistry Exam Guide

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