#### IGCSE - HONORS - ELECTROCHEMISTRY NOTES AND STUDY GUIDE:

#### 10. Electrolysis of copper - refining copper

Electrodes: metallic copper

Electrolyte: \_\_Copper sulfate Solution

Ions Present: Cu2+ SO4 2- OH- H+



In the Cathode Cu 2+ + 2e- -> Cu

In the Anode Cu -> Cu 2+ + 2e-

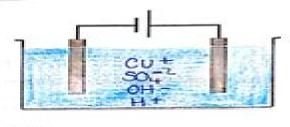
Overall Equation CU 2+ + Cu -> Cu + Cu 2+

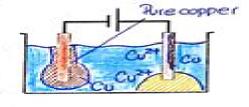
#### Observations:

Cathode: Dork solid deposit (pure copper).....

Anode: Desintegration of anode and yellow deposit below

Uses of copper\_Electrical\_wiring\_capper\_plating\_\_\_\_\_





(show in the graph how it will look like) Why is it necessary to purify Copper? In this way, it will have more conductivity.

What are the elements present in the stag? Aq., Au., Pt.

#### 11. Electroplating with Copper

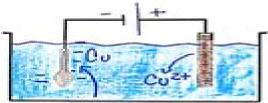
#### Electrodes:

Cathode: Spoon (metal more reactive than Cu)

Anode: Pure copper

Electrolyte: Copper Sulfate or Nitrate

lons Present: CU 21, SO 4 2-, QH -, H+



#### Half equations:

In the Cathode (002++2e -> Cu

In the Anode Cu -> Cu 2+ + 2e -

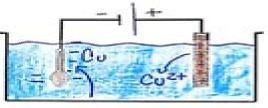
Overall Equation  $Cu^{2+} + 2e^- + Cu \rightarrow Cu + Cu^{2+}$ 

#### Observations:

Cathoda: the color of the spoon becomes douber

Anodo: Disintegrate Copper.

Artalia Sanchez



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Thirumalaisamy P. Velavancorresponding

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- 1. Understanding the eBook Matter And Change Electrochemistry Study Guide Answer
  - The Rise of Digital Reading Matter And Change Electrochemistry Study Guide Answer
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Matter And Change Electrochemistry Study Guide Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Matter And Change Electrochemistry Study Guide Answer
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Matter And Change Electrochemistry Study Guide Answer
  - Personalized Recommendations
  - Matter And Change Electrochemistry Study Guide Answer User Reviews and Ratings
  - Matter And Change Electrochemistry Study Guide Answer and Bestseller Lists

- 5. Accessing Matter And Change Electrochemistry Study Guide Answer Free and Paid eBooks
  - Matter And Change Electrochemistry Study Guide Answer Public Domain eBooks
  - Matter And Change Electrochemistry Study Guide Answer eBook Subscription Services
  - Matter And Change Electrochemistry Study Guide Answer Budget-Friendly Options
- 6. Navigating Matter And Change Electrochemistry Study Guide Answer eBook Formats
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  - Matter And Change Electrochemistry Study Guide Answer Compatibility with Devices
  - Matter And Change Electrochemistry Study Guide Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Matter And Change Electrochemistry Study Guide Answer
  - Highlighting and Note-Taking Matter And Change Electrochemistry Study Guide Answer
  - Interactive Elements Matter And Change Electrochemistry Study Guide Answer
- 8. Staying Engaged with Matter And Change Electrochemistry Study Guide Answer
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Matter And Change Electrochemistry Study Guide Answer
- 9. Balancing eBooks and Physical Books Matter And Change Electrochemistry Study Guide Answer
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Matter And Change Electrochemistry Study Guide Answer
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Matter And Change Electrochemistry Study Guide Answer
  - Setting Reading Goals Matter And Change Electrochemistry Study Guide Answer
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Matter And Change Electrochemistry Study Guide Answer
  - Fact-Checking eBook Content of Matter And Change Electrochemistry Study Guide Answer
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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