

- 1** This question is about reversible reactions.

Sulfur trioxide is produced in a reversible reaction.

The equation for the reaction is:



- 1.1** What is the correct balanced equation for the reaction?

Tick **one** box.

- ☐ $2\text{SO}_2 + \text{O}_2 \rightleftharpoons \text{SO}_3$
☐ $\text{SO}_2 + 2\text{O} \rightleftharpoons \text{SO}_3$
☐ $2\text{SO}_2 + \text{O}_2 \rightleftharpoons 2\text{SO}_3$
☐ $2\text{SO}_2 + 2\text{O}_2 \rightleftharpoons 2\text{SO}_3$

[1 mark]

- 1.2** The reaction of sulfur dioxide and oxygen to form sulfur trioxide is exothermic.

Which of the following statements is true?

Tick **one** box.

- ☐ The activation energy for the reaction of sulfur dioxide with oxygen is negative.
☐ Sulfur trioxide is colder than sulfur dioxide and oxygen.
☐ Heating the reaction will cause more sulfur trioxide to form.
☐ The reaction of sulfur trioxide to form sulfur dioxide and oxygen is endothermic.

[1 mark]

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JE Gale



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Industrial Chemistry Mark Anthony Benvenuto, 2013-12-12 Industrial Chemistry is a book that brings the subject matter of a chemistry curriculum to life. Comprehensibly written, it examines the major chemistry performed by industry and looks at how such chemical processes affect our lives. In addition, as each process is presented and examined, there is a significant discussion dedicated to the by-products, pollution, necessary waste generated, and attempts to make each process ecologically friendlier or greener. It bridges the divide between the basic chemistry that students learn in their undergraduate curriculum and the broader chemical processes that are used in real life.

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