



**55. Phosphorus has three allotropic forms — (i) white phosphorus (ii) red phosphorus and (iii) black phosphorus. Write the difference between white and red phosphorus on the basis of their structure and reactivity.**

**Solution:**

1. White phosphorus

(i) It is soft waxy solid having garlic odor

(ii) It is poisonous

(iii) Has low melting point and boiling point because  $P_4$  molecules are held together by weak van der Waal forces of attraction

2. Red Phosphorus

(i) It is hard, crystalline, odourless solid

(ii) It is non-poisonous

(iii) It has a high melting point because of its polymeric structure.

**56. Give an example to show the effect of concentration of nitric acid on the formation of oxidation product.**

**Solution:**



**57.  $PCl_5$  reacts with finely divided silver on heating and a white silver salt is obtained, which dissolves on adding excess aqueous  $NH_3$  solution. Write the reactions involved to explain what happens.**

**Solution:**

$PCl_5$  reacts with silver to form white silver salt ( $AgCl$ ). Which then dissolves on adding excess aqueous  $NH_3$  to form a soluble complex



**58. Phosphorus forms a number of oxoacids. Out of these oxoacids phosphinic acid has strong**

# Glencoe Chemistry Chapter 19 Assessment Answers

**CL Gary**



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