

## UNIT REVIEW - BONDING

### Intramolecular Forces

1. Define electronegativity

- the tendency of atoms to attract  $e^-$

2. Electronegativity increases as you go across the periodic table, and decreases as you go down a group.

3. Explain why atoms with smaller atomic radii have stronger electronegativities

- smaller atoms means  $e^-$ 's are closer to the nucleus, which has the protons
- if you decrease the distance between protons +  $e^-$ 's, you increase attractive forces.

4. Define intramolecular force

- forces that hold atoms together: ionic, covalent, polar covalent bond

5. Compare and contrast covalent, polar covalent, and ionic bond.

Covalent	Polar covalent	Ionic
<ul style="list-style-type: none"> <li>- <math>e^-</math> are shared equally</li> <li>- between 2 non-metals</li> <li>- difference in electronegativity is <math>\leq 0.2</math></li> <li>- atoms are neutral</li> </ul>	<ul style="list-style-type: none"> <li>- <math>e^-</math> are shared unequally</li> <li>- between 2 non-metals</li> <li>- difference in electronegativity is <math>0.3 - 1.6</math></li> <li>- atoms are partially charged</li> </ul>	<ul style="list-style-type: none"> <li>- <math>e^-</math> are transferred</li> <li>- between metal + non metal</li> <li>- difference in electronegativity is <math>\geq 1.7</math></li> <li>- atoms have a charge</li> </ul>

6. What is a partial charge and how is it different from an ionic charge?

A partial charge results from an unequal sharing of electrons between 2 atoms

An ionic charge results from a transfer of electrons from 1 atom to another

# Flocabulary Unit 10 Changes Answer Key

**Xiang Xie**



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## **Table of Contents Flocabulary Unit 10 Changes Answer Key**

1. Understanding the eBook Flocabulary Unit 10 Changes Answer Key
  - The Rise of Digital Reading Flocabulary Unit 10 Changes Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Flocabulary Unit 10 Changes Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Flocabulary Unit 10 Changes Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Flocabulary Unit 10 Changes Answer Key
  - Personalized Recommendations
  - Flocabulary Unit 10 Changes Answer Key User Reviews and Ratings
  - Flocabulary Unit 10 Changes Answer Key and Bestseller Lists
5. Accessing Flocabulary Unit 10 Changes Answer Key Free and Paid eBooks
  - Flocabulary Unit 10 Changes Answer Key Public Domain eBooks
  - Flocabulary Unit 10 Changes Answer Key eBook Subscription Services
  - Flocabulary Unit 10 Changes Answer Key Budget-Friendly Options
6. Navigating Flocabulary Unit 10 Changes Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Flocabulary Unit 10 Changes Answer Key Compatibility with Devices
- Flocabulary Unit 10 Changes Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Flocabulary Unit 10 Changes Answer Key
  - Highlighting and Note-Taking Flocabulary Unit 10 Changes Answer Key
  - Interactive Elements Flocabulary Unit 10 Changes Answer Key
- 8. Staying Engaged with Flocabulary Unit 10 Changes Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Flocabulary Unit 10 Changes Answer Key
- 9. Balancing eBooks and Physical Books Flocabulary Unit 10 Changes Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Flocabulary Unit 10 Changes Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Flocabulary Unit 10 Changes Answer Key
  - Setting Reading Goals Flocabulary Unit 10 Changes Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Flocabulary Unit 10 Changes Answer Key
  - Fact-Checking eBook Content of Flocabulary Unit 10 Changes Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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