

UNIT REVIEW - BONDING

Intramolecular Forces

1. Define electronegativity

- the tendency of atoms to attract e^-

2. Electronegativity increases as you go across the periodic table, and decreases as you go down a group.

3. Explain why atoms with smaller atomic radii have stronger electronegativities

- smaller atoms means e^- 's are closer to the nucleus, which has the protons
- if you decrease the distance between protons + e^- 's, you increase attractive forces.

4. Define intramolecular force

- forces that hold atoms together: ionic, covalent, polar covalent bond

5. Compare and contrast covalent, polar covalent, and ionic bond.

Covalent	Polar covalent	Ionic
<ul style="list-style-type: none"> - e^- are shared equally - between 2 non-metals - difference in electronegativity is ≤ 0.2 - atoms are neutral 	<ul style="list-style-type: none"> - e^- are shared unequally - between 2 non-metals - difference in electronegativity is $0.3 - 1.6$ - atoms are partially charged 	<ul style="list-style-type: none"> - e^- are transferred - between metal + non metal - difference in electronegativity is ≥ 1.7 - atoms have a charge

6. What is a partial charge and how is it different from an ionic charge?

A partial charge results from an unequal sharing of electrons between 2 atoms
 An ionic charge results from a transfer of electrons from 1 atom to another

Flocabulary Transformation Answer Key Unit

Michael Rajnik



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