

F321: Atoms, Bonds and Groups

Redox

39 Marks

1. The element strontium forms a nitrate, $\text{Sr}(\text{NO}_3)_2$, which decomposes on heating as shown below.



- (i) Using oxidation numbers, explain why the reaction involves both oxidation and reduction.

.....

.....

.....

.....

[3]

- (ii) A student heats 5.29 g of $\text{Sr}(\text{NO}_3)_2$ and collects the gas at room temperature and pressure, RTP.



Calculate the volume of gas, in dm^3 , obtained by the student at RTP.

Molar mass of $\text{Sr}(\text{NO}_3)_2 = 211.6 \text{ g mol}^{-1}$.

answer = dm^3

[3]

[Total 6 marks]

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Table of Contents F321 Atoms Bonds And Groups June 2014

1. Understanding the eBook F321 Atoms Bonds And Groups June 2014
 - The Rise of Digital Reading F321 Atoms Bonds And Groups June 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying F321 Atoms Bonds And Groups June 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an F321 Atoms Bonds And Groups June 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from F321 Atoms Bonds And Groups June 2014
 - Personalized Recommendations
 - F321 Atoms Bonds And Groups June 2014 User Reviews and Ratings
 - F321 Atoms Bonds And Groups June 2014 and Bestseller Lists
5. Accessing F321 Atoms Bonds And Groups June 2014 Free and Paid eBooks
 - F321 Atoms Bonds And Groups June 2014 Public Domain eBooks
 - F321 Atoms Bonds And Groups June 2014 eBook Subscription Services
 - F321 Atoms Bonds And Groups June 2014 Budget-Friendly Options

6. Navigating F321 Atoms Bonds And Groups June 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - F321 Atoms Bonds And Groups June 2014 Compatibility with Devices
 - F321 Atoms Bonds And Groups June 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of F321 Atoms Bonds And Groups June 2014
 - Highlighting and Note-Taking F321 Atoms Bonds And Groups June 2014
 - Interactive Elements F321 Atoms Bonds And Groups June 2014
8. Staying Engaged with F321 Atoms Bonds And Groups June 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers F321 Atoms Bonds And Groups June 2014
9. Balancing eBooks and Physical Books F321 Atoms Bonds And Groups June 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection F321 Atoms Bonds And Groups June 2014
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine F321 Atoms Bonds And Groups June 2014
 - Setting Reading Goals F321 Atoms Bonds And Groups June 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of F321 Atoms Bonds And Groups June 2014
 - Fact-Checking eBook Content of F321 Atoms Bonds And Groups June 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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