2. Magnesium carbonate reacts with hydrochloric acid.

$$MqCO_{s}(s) + 2HCl(aq) \rightarrow MgCl_{s}(aq) + CO_{s}(q) + H_{s}O(l)$$

(a) What mass of magnesium carbonate would react with excess hydrochloric acid to produce 240 cm² of carbon dioxide, measured at room temperature and pressure?

Data: I mol of any gas occupies 24.0 dm1 at room temperature and pressure

11.74	ita: I mol of any gas occupies 24.0 dm1 at re	som temperature and pressure
	Molar mass of magnesium carbonate =	an Diversity O. L. Holling
DA A	oner of	CO = Volume in dm3 = 0.24 = 0.01
CLB	8.430	
		ontains same moles as CD2
□ D	843g moles = m	1055; mass=moles x Mr
(b) W/	ray may not 0 250 modelm. I hwelenethloric acid?	bonate needed to neutralise = 0.843g
E A	0.423g Concentration of HCI:	volindm3 MgCO3:HCI
DG B		
DC		
$\Box$ D	= 0.0125m	10/es of HC1 = x0.0125: 0.0125
XI A	A colourless solution - all Myco3	s unique
□ 8	A coloured solution and colo	MgCO3 HCI
	A white precipitate Mace	Colorless moles moles
III D		mass = moles x Mr
		(Total for Question 2 = 3 marks)
		$= 6.25 \times 10 \times 84.3$
	D   C   D   C   D   C   W   X   A   D   C   C   C   C   C   C   C   C   C	□ B 8.419 □ C 84.39 □ D 8439 □ D 8439 □ Mg CO3 C moles = m  (b) What is the minimum mass of magnesium carl  50.0 cm² of 0.250 moldm ¹ hydrochloric acid? □ A 0.4239 □ Concentration of HCI: □ A 0.4239 □ Concentration of HCI: □ C 1.059 □ C 1.059 □ C 0.250 X □ D 2.119 □ C 0.0125m  (c) What would be seen at the end of the reaction  X A A colourless solution → all MgCO3 □ B A coloured solution and colour

# **Multiple Choice Questions On Stoichiometry**

**Léo Pomar** 

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