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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al $\times \frac{1 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \times \frac{101.964 \text{ g Al}_2\text{O}_3}{302.3 \text{ g Al}_2\text{O}_3} = 26.982 \text{ g Al}_2\text{O}_3$ c determine the percent yield $\frac{26.0 \text{ g}}{26.982 \text{ g}} \times 100 = 96.4\%$

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 $\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100$