

Experiment 6B: Calculations with a Chemical Reaction

Objectives:

1. To understand the theory of stoichiometry, which is an essential concept and tool for chemists, for quantitative analysis.
2. To calculate the limiting reagent, which is the reagent present in the smallest amount.
3. To calculate the theoretical yield, which is the maximum amount of product that can be formed.
4. To calculate the percent yield, which is the ratio of the actual yield to the theoretical yield.
5. To calculate the percent error, which is the ratio of the absolute error to the theoretical yield.

Materials: Please refer to Experiment 6B page 95 of the Pharmacokinetics Laboratory.

Procedure:

Preparation:

1. Prepare a 100 mL solution of 0.1 M sodium hydroxide.
2. Prepare a 100 mL solution of 0.1 M hydrochloric acid.
3. Prepare 10 mL of 0.1 M sodium hydroxide solution and 10 mL of 0.1 M hydrochloric acid.
4. Prepare 10 mL of 0.1 M sodium hydroxide solution and 10 mL of 0.1 M hydrochloric acid.
5. Prepare 10 mL of 0.1 M sodium hydroxide solution and 10 mL of 0.1 M hydrochloric acid.
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9. Prepare 10 mL of 0.1 M sodium hydroxide solution and 10 mL of 0.1 M hydrochloric acid.
10. Prepare 10 mL of 0.1 M sodium hydroxide solution and 10 mL of 0.1 M hydrochloric acid.

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M Walker

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