

**EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES
(CHEMISTRY)**

TABLE 1 PHYSICAL CONSTANTS

NAME	SYMBOL	VALUE
Magnitude of charge on electron	e	$1,6 \times 10^{-19} \text{ C}$
Mass of an electron	m_e	$9,1 \times 10^{-31} \text{ kg}$
Standard pressure	p°	$1,01 \times 10^5 \text{ Pa}$
Molar gas volume at STP	V_m	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature	T°	273 K
Avogadro's constant	N_A	$6,02 \times 10^{23} \text{ mol}^{-1}$
Faraday's constant	F	$96\,500 \text{ C} \cdot \text{mol}^{-1}$

TABLE 2 CHEMISTRY FORMULAE

$n = \frac{m}{M}$	$n = \frac{N}{N_A}$	$n = \frac{V}{V_m}$
$c = \frac{n}{V}$ OR $c = \frac{m}{MV}$	$K_w = [\text{H}_3\text{O}^+] \cdot [\text{OH}^-] = 1 \times 10^{-14}$ at 25 °C (298 K)	
$q = It$ $q = nF$	$E_{\text{cell}}^\circ = E_{\text{cathode}}^\circ - E_{\text{anode}}^\circ$ $E_{\text{cell}}^\circ = E_{\text{oxidising agent}}^\circ - E_{\text{reducing agent}}^\circ$	

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1. Understanding the eBook Ieb Grade 1physical Science Past Papers
 - The Rise of Digital Reading Ieb Grade 1physical Science Past Papers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ieb Grade 1physical Science Past Papers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ieb Grade 1physical Science Past Papers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ieb Grade 1physical Science Past Papers
 - Personalized Recommendations
 - Ieb Grade 1physical Science Past Papers User Reviews and Ratings
 - Ieb Grade 1physical Science Past Papers and Bestseller Lists
5. Accessing Ieb Grade 1physical Science Past Papers Free and Paid eBooks
 - Ieb Grade 1physical Science Past Papers Public Domain eBooks
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 - Ieb Grade 1physical Science Past Papers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Ieb Grade 1physical Science Past Papers Compatibility with Devices
 - Ieb Grade 1physical Science Past Papers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ieb Grade 1physical Science Past Papers
 - Highlighting and Note-Taking Ieb Grade 1physical Science Past Papers
 - Interactive Elements Ieb Grade 1physical Science Past Papers
8. Staying Engaged with Ieb Grade 1physical Science Past Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ieb Grade 1physical Science Past Papers
9. Balancing eBooks and Physical Books Ieb Grade 1physical Science Past Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ieb Grade 1physical Science Past Papers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ieb Grade 1physical Science Past Papers
 - Setting Reading Goals Ieb Grade 1physical Science Past Papers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ieb Grade 1physical Science Past Papers
 - Fact-Checking eBook Content of Ieb Grade 1physical Science Past Papers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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