

Worksheet: Writing and Balancing Chemical Reactions

1. Balance the following equations and indicate the type of reaction as formation, decomposition, single replacement, double replacement, hydrocarbon combustion, or other.

- $\text{Cu}_{(s)} + \text{O}_{2(g)} \rightarrow \text{CuO}_{(s)}$
- $\text{H}_2\text{O}_{(g)} \rightarrow \text{H}_{2(g)} + \text{O}_{2(g)}$
- $\text{Fe}_{(s)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_{2(g)} + \text{Fe}_3\text{O}_{4(s)}$
- $\text{AsCl}_{3(aq)} + \text{H}_2\text{S}_{(aq)} \rightarrow \text{As}_2\text{S}_{3(s)} + \text{HCl}_{(aq)}$
- $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}_{(s)} \rightarrow \text{CuSO}_{4(s)} + \text{H}_2\text{O}_{(g)}$
- $\text{Fe}_2\text{O}_{3(s)} + \text{H}_2_{(g)} \rightarrow \text{Fe}_{(s)} + \text{H}_2\text{O}_{(l)}$
- $\text{CaCO}_{3(s)} \rightarrow \text{CaO}_{(s)} + \text{CO}_{2(g)}$
- $\text{Fe}_{(s)} + \text{S}_{8(s)} \rightarrow \text{FeS}_{(s)}$
- $\text{H}_2\text{S}_{(aq)} + \text{KOH}_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{K}_2\text{S}_{(aq)}$
- $\text{NaCl}_{(s)} \rightarrow \text{Na}_{(s)} + \text{Cl}_{2(g)}$
- $\text{Al}_{(s)} + \text{H}_2\text{SO}_{4(aq)} \rightarrow \text{H}_{2(g)} + \text{Al}_2(\text{SO}_4)_{3(aq)}$
- $\text{H}_3\text{PO}_4(aq) + \text{NH}_4\text{OH}_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)} + (\text{NH}_4)_3\text{PO}_4(aq)$
- $\text{C}_3\text{H}_8_{(g)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- $\text{Al}_{(s)} + \text{O}_{2(g)} \rightarrow \text{Al}_2\text{O}_{3(s)}$
- $\text{CH}_4_{(g)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- $\text{K}_2\text{SO}_{4(aq)} + \text{BaCl}_{2(aq)} \rightarrow \text{KCl}_{(aq)} + \text{BaSO}_{4(s)}$
- $\text{C}_5\text{H}_{12(l)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$
- $\text{Ca}(\text{OH})_{2(aq)} + \text{NH}_4\text{Cl}_{(aq)} \rightarrow \text{NH}_4\text{OH}_{(aq)} + \text{CaCl}_{2(aq)}$
- $\text{V}_2\text{O}_5_{(s)} + \text{Ca}_{(s)} \rightarrow \text{CaO}_{(s)} + \text{V}_{(s)}$
- $\text{Na}_{(s)} + \text{ZnI}_{2(aq)} \rightarrow \text{NaI}_{(aq)} + \text{Zn}_{(s)}$
- $\text{C}_7\text{H}_6\text{O}_{3(l)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- $\text{Ca}_{(s)} + \text{N}_{2(g)} \rightarrow \text{Ca}_3\text{N}_{2(s)}$
- $\text{Fe}_2\text{O}_{3(s)} + \text{H}_{2(g)} \rightarrow \text{Fe}_{(s)} + \text{H}_2\text{O}_{(l)}$
- $\text{C}_{15}\text{H}_{30(l)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$
- $\text{BN}_{(s)} + \text{F}_{2(g)} \rightarrow \text{BF}_{3(s)} + \text{N}_{2(g)}$
- $\text{C}_{12}\text{H}_{26(l)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$

2. Predict the product(s) along with the states, indicate the type of reaction, and balance the following chemical reactions.

- A solution of lead (II) nitrate is mixed with a solution of sodium iodide.
- Solid zinc sulfide reacts with oxygen in the air.
- Liquid butane ($\text{C}_4\text{H}_{10(l)}$) is used as a fuel to ignite a lighter.
- Barium hydroxide solution is neutralized by adding hydrochloric acid ($\text{HCl}_{(aq)}$).
- Copper metal is placed in a solution of silver nitrate.
- Sulfur burns in oxygen to make sulfur dioxide gas.
- A solution of aluminum sulfate is mixed with a solution of calcium hydroxide.
- Zinc metal is placed in sulfuric acid ($\text{H}_2\text{SO}_{4(aq)}$).
- Aluminum powder is placed in a container filled with chlorine gas.
- Sucrose undergoes cellular respiration.

Identifying And Balancing Chemical Equations Answers

Chapter 8

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Table of Contents Identifying And Balancing Chemical Equations Answers Chapter 8

1. Understanding the eBook Identifying And Balancing Chemical Equations Answers Chapter 8
 - The Rise of Digital Reading Identifying And Balancing Chemical Equations Answers Chapter 8
 - Advantages of eBooks Over Traditional Books
2. Identifying Identifying And Balancing Chemical Equations Answers Chapter 8
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Identifying And Balancing Chemical Equations Answers Chapter 8
 - User-Friendly Interface
4. Exploring eBook Recommendations from Identifying And Balancing Chemical Equations Answers Chapter 8
 - Personalized Recommendations
 - Identifying And Balancing Chemical Equations Answers Chapter 8 User Reviews and Ratings

- Identifying And Balancing Chemical Equations Answers Chapter 8 and Bestseller Lists
- 5. Accessing Identifying And Balancing Chemical Equations Answers Chapter 8 Free and Paid eBooks
 - Identifying And Balancing Chemical Equations Answers Chapter 8 Public Domain eBooks
 - Identifying And Balancing Chemical Equations Answers Chapter 8 eBook Subscription Services
 - Identifying And Balancing Chemical Equations Answers Chapter 8 Budget-Friendly Options
- 6. Navigating Identifying And Balancing Chemical Equations Answers Chapter 8 eBook Formats
 - ePub, PDF, MOBI, and More
 - Identifying And Balancing Chemical Equations Answers Chapter 8 Compatibility with Devices
 - Identifying And Balancing Chemical Equations Answers Chapter 8 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Identifying And Balancing Chemical Equations Answers Chapter 8
 - Highlighting and Note-Taking Identifying And Balancing Chemical Equations Answers Chapter 8
 - Interactive Elements Identifying And Balancing Chemical Equations Answers Chapter 8
- 8. Staying Engaged with Identifying And Balancing Chemical Equations Answers Chapter 8
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Identifying And Balancing Chemical Equations Answers Chapter 8
- 9. Balancing eBooks and Physical Books Identifying And Balancing Chemical Equations Answers Chapter 8
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Identifying And Balancing Chemical Equations Answers Chapter 8
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Identifying And Balancing Chemical Equations Answers Chapter 8
 - Setting Reading Goals Identifying And Balancing Chemical Equations Answers Chapter 8
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Identifying And Balancing Chemical Equations Answers Chapter 8
 - Fact-Checking eBook Content of Identifying And Balancing Chemical Equations Answers Chapter 8
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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