Name Period Date

8 • Why Do Hot Air Balloons Float?

IDEAL

PV = nRT where

P = pressure in atmosphere

V = volume in liters

n = number of moles of gas

R = Universal Gas Constant = 0.0821 L atm/mol·K

T = Kelvin temperature

How many moles of oxygen will occupy a volume of 2.50 liters at 1.20 atm and 25°C?

PU=nRT

W-3-1- (242) L

40%

8.

What volume will 2.00 moles of nitrogen occupy at 720; mmHg and 20.°C? 2.

. 568 mul

What pressure will be exerted by 25.0 g of CO2 at temperature of 25 °C and a volume of 500 mL? 3... ary a glank

. 0704 me

- At what temperature will 5.00 g of Cb exert a pressure of 200 mmHg at a volume of 750; mL?
- 1.05 at 298K What is the density of NH at 800, mmHe and 25 °C? 5...

.980 at 293 K

- If the density of a gas is 1.2 g/L at 745 mmHg and 20.0°C, what is its molar mass?
- 547 L 8.79 atm How many moles of nitrogen gas will occupy a volume of 347 mt. at 6680 mmHg and 27 °C? T_{c}

225 ---

218K

- What volume will 454 grams (1 lb) of hydrogen occupy at 1.05 atm and 25 °C?
- 6.03 ch Find the number of grams of CO2 that exert a pressure of 285-mmHg at a volume of 32.5 L and a temperature of 32°C. 3.05 K

An elemental gas has a mass of 10.3 g. If the volume is 58.4 L and the pressure is 758-mmHg at a 10. temperature of 2.5°C, what is the gas? FIND MOLON my 5 3

275.5 K

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