

## Notes:

- i. All questions are compulsory.
- ii. Answers to the two sections are to be written in the same answer book.
- iii. Figures to the right hand side indicate full marks.
- iv. Write balanced chemical equations and draw neat and labelled diagrams wherever necessary.
- v. Every new question must be started on a new page.
- vi. Use of logarithmic table is allowed.

## CHEMISTRY: SECTION – I

Q. 1. Select and write the most appropriate answer from the given alternatives for each sub-question: [7]

- i. In body centred cubic structure the space occupied is about  
(A) 68 % (B) 53 %  
(C) 38 % (D) 32 %
- ii. For a gaseous reaction the unit of rate of reaction is  
(A)  $\text{L atm s}^{-1}$  (B)  $\text{atm mol}^{-1} \text{s}^{-1}$   
(C)  $\text{atm s}^{-1}$  (D)  $\text{mol s}$
- iii. Which of the following compounds contains S – O as well as S – S bonds?  
(A) Sulphuric acid (B) Thiosulphuric acid  
(C) Sulphurous acid (D) Thiosulphurous acid
- iv. Which of the following solutions shows maximum depression in freezing point?  
(A) 0.5 M  $\text{Li}_2\text{SO}_4$  (B) 1 M NaCl  
(C) 0.5 M  $\text{Al}_2(\text{SO}_4)_3$  (D) 0.5 M  $\text{BaCl}_2$
- v. For a chemical reaction  $\Delta S = -0.035 \text{ kJ/K}$  and  $\Delta H = -20 \text{ kJ}$ .  
(A) 5.14 K (B) 57.14 K  
(C) 571.4 K (D) 5714.0 K
- vi. The standard e.m.f of the following cell is 0.463 V  
 $\text{Cu} | \text{Cu}_{(0.50)}^{++} || \text{Ag}_{(0.50)}^+ | \text{Ag}$ . If  $E_{\text{Ag}}^{\circ} = 0.800 \text{ V}$ ,  
What is the standard potential of Cu electrode?  
(A) 1.137 V (B) 0.337 V  
(C) 0.463 V (D) -0.463 V
- vii.  $\text{Fe}_2\text{O}_3$  is reduced to spongy iron near the top of blast furnace by  
(A)  $\text{H}_2$  (B) CaO  
(C)  $\text{SiO}_2$  (D) CO

# Hsc Board Chemistry Paper Of 2013 Solution

**Peter Lewis, Roger Michael Slade**



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