

Name Rocio Castro Period 4

Chapter 3: Water and the Fitness of the Environment

Concept 3.1 The polarity of water molecules results in hydrogen bonding

- Study the water molecules at the right. On the central molecule, label oxygen (O) and hydrogen (H).
- What is a *polar molecule*? Why is water considered polar?

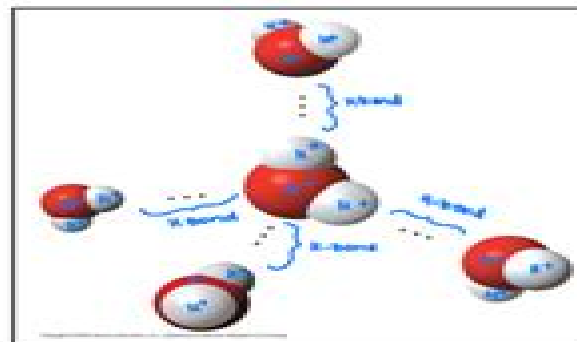
Polar molecule
-unequal sharing of electrons
-charge unevenly distributed

oxygen has 6 e⁻ and 2 protons
H = +1 O = -1
 $2 \times +1 + 1 \times -1 = -2 + 1 = -1$

- Now, add + and - signs to indicate the charged regions of each molecule. Then, indicate the hydrogen bonds.
- Explain *hydrogen bonding*. How many hydrogen bonds can a single water molecule form?

4 H-bonds max per water molecule

H-bonding: noncovalent attraction between hydrogen & an electronegative atom nearby



Concept 3.2 Four emergent properties of water contribute to Earth's fitness for life

Hydrogen bonding accounts for the unique properties of water. Let's look at several.

Cohesion

- Distinguish between *cohesion* and *adhesion*.
Cohesion: the H-bond tension between water molecules (intermolecular forces)
Adhesion: the clinging of one substance to another (bonding between 2 unlike molecules) water H₂O on side of cup after finish drinking
- What is demonstrated when you see beads of water on a waxed car hood?
cohesion & surface tension
- Which property explains the ability of a water strider to walk on water?
Surface tension

Moderation of Temperature

- The calorie is a unit of heat. Define *calorie*.

The amount of heat it takes to raise the temp of 1g of water by 1°

- Water has high *specific heat*. What does this mean? How does water's specific heat compare to alcohol's?
specific heat is the amount of heat that causes 1g of water to change temp by 1°C. water specific heat is 1 cal/gc. Alcohol requires less heat to change temp since it's specific heat is 0.69 cal/gc.
 - Explain how hydrogen bonding contributes to water's high specific heat.
Hydrogen bonds are all over water molecules & it takes a lot of heat to break the H-bonds and a good amount of heat is released when H-bonds form
- * Because H-bonds, water molecules move faster

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