PHASES OF MATTER

Fame hey	

Date Period

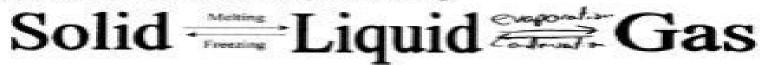
Phase Chariges

Heating a substance is a given phase generally causes the temperature to increase. This causes the particles to move flaster and coilide harder. Harder, firster collisions cause particles to rebound harder moving them flarther apart. Larger distances between particles weakers the forces of attraction however them. When the forces of attraction are weak enough, the distance between the particles increases markedly and the phase changes. As a result, a solid melts and a liquid evaporates. The reverse happens when a substance cools. As the particles slow down and the collisions weakers, the particles move closer together increasing the forces of attraction between the particles. As a result, a gas condenses and a liquid fractor. At normal immospheric pressures, some substances change directly from a solid to a gas. This change is coilled subtimation. This is what occurs when we cabes strink in the fractor, or when a wash dries on a clothesing on a cold winter day. Some substances, such as day ice [CO₂(x)], never pass through a liquid phase at mandard pressure [101.3 kPa]. A gas can also change directly to a solid. This is called deposition.



Answer the questions below based on your reading and on your knowledge of chemistry.

In the space below are the labels sollid, fiquid, and gas. Draw arrows between them to represent phase changes. Label each
arrow with the name of the phase change it represents. Two examples are done for you. Also, draw an arrow labeled
temperature to show the direction in which temperature is increasing.



- 2. Under what conditions does a mote many ways When Kinetic energy (temperature)

 Yister a have the forces of attraction, a solid milt
- decreases and the forces of attraction are foll by the fartiels

Evan Silberstein Chemistry Answer Key

Emilie Sanchez

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Table of Contents Evan Silberstein Chemistry Answer Key

- 1. Understanding the eBook Evan Silberstein Chemistry Answer Key
 - The Rise of Digital Reading Evan Silberstein Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Evan Silberstein Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Evan Silberstein Chemistry Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Evan Silberstein Chemistry Answer Key
 - Personalized Recommendations
 - Evan Silberstein Chemistry Answer Key User Reviews and Ratings
 - Evan Silberstein Chemistry Answer Key and Bestseller Lists
- 5. Accessing Evan Silberstein Chemistry Answer Key Free and Paid eBooks

- Evan Silberstein Chemistry Answer Key Public Domain eBooks
- Evan Silberstein Chemistry Answer Key eBook Subscription Services
- Evan Silberstein Chemistry Answer Key Budget-Friendly Options
- 6. Navigating Evan Silberstein Chemistry Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Evan Silberstein Chemistry Answer Key Compatibility with Devices
 - Evan Silberstein Chemistry Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Evan Silberstein Chemistry Answer Key
 - Highlighting and Note-Taking Evan Silberstein Chemistry Answer Key
 - Interactive Elements Evan Silberstein Chemistry Answer Key
- 8. Staying Engaged with Evan Silberstein Chemistry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Evan Silberstein Chemistry Answer Key
- 9. Balancing eBooks and Physical Books Evan Silberstein Chemistry Answer Key
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Evan Silberstein Chemistry Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Evan Silberstein Chemistry Answer Key
 - Setting Reading Goals Evan Silberstein Chemistry Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Evan Silberstein Chemistry Answer Key
 - Fact-Checking eBook Content of Evan Silberstein Chemistry Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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