

2. Magnesium carbonate reacts with hydrochloric acid.



- (a) What mass of magnesium carbonate would react with excess hydrochloric acid to produce 240cm^3 of carbon dioxide, measured at room temperature and pressure?

Data: 1 mol of any gas occupies 24.0dm^3 at room temperature and pressure

Molar mass of magnesium carbonate = 84.3g mol^{-1}

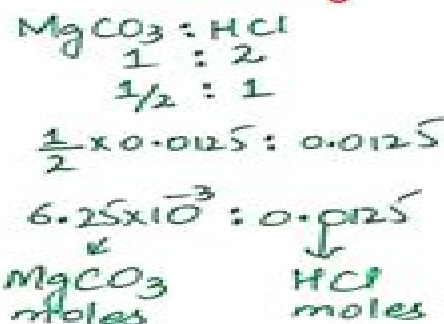
- ☒ A 0.843 g
☐ B 8.43 g
☐ C 84.3 g
☐ D 843 g

$$\begin{aligned}\text{Volume of CO}_2 &= 240\text{cm}^3 = 0.24\text{dm}^3 \\ \text{moles of CO}_2 &= \frac{\text{Volume in dm}^3}{24} = \frac{0.24}{24} = 0.01 \\ \text{MgCO}_3 &\text{ contains same moles as CO}_2 \\ \text{moles} &= \frac{\text{mass}}{\text{Mr}} ; \text{ mass} = \text{moles} \times \text{Mr} \\ &= 0.01 \times 84.3 \\ &= 0.843\text{g}\end{aligned}$$

- (b) What is the minimum mass of magnesium carbonate needed to neutralise 50.0cm^3 of 0.250mol dm^{-3} hydrochloric acid?

- ☐ A 0.423 g
☒ B 0.527 g
☐ C 1.05 g
☐ D 2.11 g

$$\begin{aligned}\text{Concentration of HCl} &= \frac{\text{moles}}{\text{Vol in dm}^3} \\ \text{moles} &= \text{Conc} \times \text{Vol in dm}^3 \\ &= 0.250 \times 0.05 \\ &= 0.0125\text{moles of HCl}\end{aligned}$$



- (c) What would be seen at the end of the reaction with excess acid?

- ☒ A A colourless solution
☐ B A coloured solution
☐ C A white precipitate
☐ D A coloured precipitate

all MgCO_3 is utilized
and color of
 MgCO_2 colorless

$$\text{mass} = \text{moles} \times \text{Mr}$$

(Total for Question 2 = 3 marks)

$$\begin{aligned}&= 6.25 \times 10^{-3} \times 84.3 \\ &= 0.527\text{g}\end{aligned}$$

Multiple Choice Questions Of Stoichiometry

Arshad Iqbal



Multiple Choice Questions Of Stoichiometry:

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This book delves into Multiple Choice Questions Of Stoichiometry. Multiple Choice Questions Of Stoichiometry is a vital topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Multiple Choice Questions Of Stoichiometry, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Multiple Choice Questions Of Stoichiometry
- Chapter 2: Essential Elements of Multiple Choice Questions Of Stoichiometry
- Chapter 3: Multiple Choice Questions Of Stoichiometry in Everyday Life
- Chapter 4: Multiple Choice Questions Of Stoichiometry in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Multiple Choice Questions Of Stoichiometry. This chapter will explore what Multiple Choice Questions Of Stoichiometry is, why Multiple Choice Questions Of Stoichiometry is vital, and how to effectively learn about Multiple Choice Questions Of Stoichiometry.
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5. In chapter 4, the author will scrutinize the relevance of Multiple Choice Questions Of Stoichiometry in specific contexts. This chapter will explore how Multiple Choice Questions Of Stoichiometry is applied in specialized fields, such as education, business, and technology.
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