CHEMISTRY 10TH - Multiple Choice Questions



CHAPTER# 01

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- 01. Which one of the following statements is false about dynamic equilibrium?
 - (a) It takes place in a closed container
 - (b) Concentration of reactant and products are not changed
 - (c) Rate of the forward reaction is equal to the rate of the reverse reaction
 - (d) ▼ Equilibrium cannot be disturbed by any external stress
- 02. Consider the following reaction and indicate which of the following best describes equilibrium constant expression K_c . $4NH_3 + 5O_2 = 4NO + 6H_2O$

(a)
$$V K_c = \frac{[NO]^4 [H_2O]^6}{[NH_3]^4 [O_2]^5}$$

(c)
$$K_c = \frac{[NH_3][O_2]}{[H_1O][NO]}$$

(b)
$$K_c = \frac{[NH_3]^4 [O_2]^5}{[NO]^4 [H_2O]^6}$$

(d)
$$K_c = \frac{[4NO][6H_2O]}{[4NH_3][5O_2]}$$

- O3. A reaction which is never goes to completion is known as reversible reaction. Reversible reaction is represented by:
 - (a) doted lines

(b) single arrow

(c) V double arrow

- (d) double straight line
- 04. When the magnitude of K_c is small, indicates:
 - (a) V Reaction mixture contain most of the reactant
 - (b) Reaction mixture contain most of the product
 - (c) Reaction mixture contain almost equal amount of reactant and product
 - (d) Reaction goes to completion
- 05. For which system does the equilibrium constant, K_c has units of concentration?

(b)
$$N_{2(g)} + O_{2(g)} \Leftrightarrow 2NO_{(g)}$$

(c)
$$H_{2(g)} + I_{2(g)} \Leftrightarrow 2HI_{(g)}$$

(d)
$$CO_{2(g)} + H_{2(g)} \Rightarrow CO_{(g)} + H_2O_{(1)}$$

<u>Multiple Choice Questions And Answers In Chemical</u> <u>Equilibrium</u>

E Durkheim

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