ACIDS AND BASES QUIZ

1. Which of the following indicates an ocidic solution? A) pH = 7.5B) pH = 7C) pH = 6.5D) more than one correct response E) no correct response 2. Which of the following equations represents an acid-base neutralization reaction? B) HNO₁ + KOH → KNO₂ + H₂O A) $H_iSO_i + Zn \rightarrow ZnSO_i + H_i$ C) Ba(OH), + Na,SO, → BaSO, + 2NaOH D) more than one correct response E) none of the above 3. In which of the following pairs of substances are both members of the pair salts? A) NaCl and NaOH. B) KNO₃ and K₃SO₃ C) H₂PO₄ and (NH₄)₂PO₄ D) more than one correct response E) none of the above If the pH of a solution increases from 4.0 to 6.0, the hydronium ion concentration. A) increases by a factor of 1.5. B) increases by a factor of 100. C) decreases by a factor of 1.5. D) decreases by a factor of 100. 5. Which of the following statements concerning Arrhenius acids and Arrhenius bases is Incorrect? A) In the pure state, Arrhenius acids are covalent compounds. B) In the pure state, Arrhenius bases are ionic compounds. C) Dissociation is the process by which Arrhenius acids produce H' ions in solution. D) Arrhenius bases are also called oxides. 6. Highly basic solutions are: A) safe to handle B) sweet tasting C) corrosive and reactive like highly acidic solutions D) comosive but not reactive: When dissolved in water, sodium hydroxide (NaOH): A) turns phenolphthalein colourless.
 B) produces a high concentration of hydronium ions. C) releases O2 ions. D) produces a high concentration of hydroxide ions. E) reacts with carbonates to produce a salt and carbon dioxide gas. B. Bases taste: A) Salty (C) Sweet D) Sour B) Bitter 9. The characteristic ion of acids is: A) carbonate ion CO₁²
 B) phosphate ion PO₂³
 C) hydrogen ion (H')
 D) hydroxide ion (OH') 10. Which of the following pH values indicates the strongest base?

ID0: 1

A3 4

B0 7

11. A positive test for a base occurs when:

CB 111

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