

Test: Tuesday 2/18/20

*This guide is not the only thing you should use to study. It does not provide you with everything you need. You should also rely on your **Chemical Interactions** textbook, homework, and classroom notes. Use everything you can for the best results.*

*****Study your Solubility and Solutions Quiz***

Topic 1: Understanding Solutions – Textbook pg. 84-89; Binder pg 6-8

1. What happens to the particles of a solid when dissolved in water? What are the parts of that solution?
 - Particles of the solid leave each other and become surrounded by particles of water.
 - Ionic solids in water can conduct electricity; Molecular solids cannot!
 - Parts of the solution: Solvent:
 - Solvent = the largest amount; does the dissolving
 - Solute = the smaller amount; gets dissolved
2. Compare and contrast solutions, colloids, and suspensions.

3. Give an example of each (solutions, colloids, suspensions).

- Solutions = Air; saltwater
- Colloids = Fog; milk; Gelatin
- Solutions = Oil and Vinegar; Pepper/Water

4. How do solutes affect a solvent when added? **Explain.** (3 ways)

- a. Lower Freezing Point – Makes it harder for water molecules to form crystals of solid ice
- b. Increase Boiling Point – Makes it harder for water molecules to escape
- c. Ionic Solutes are able to conduct electricity in water – negative and positive ions are attracted to water molecules

Contrast		
Solutions	Colloids	Suspensions
*Same properties throughout *Molecules are too small to be seen **Can be made from solids, liquids, or gases	*Larger particles than a solution but still too small to be seen. **Able to scatter a beam of light	Particles can be seen and easily separated by settling or filtration DOES NOT have the same properties throughout

Topic 2: Concentration and Solubility – Textbook pg. 92-97; Binder pg 9-12

5. What is the difference between a saturated and an unsaturated solution? Give examples.
 - Saturated - add so much solute that no more dissolves
 - Unsaturated – Able to dissolve more solute
6. Explain the difference between a dilute and a concentrated solution. How is this different from a saturated and unsaturated solution?
 - Dilute - Mixture that has only a little solute dissolved in a solvent
 - Concentrated – Has a lot of solute dissolved in the solvent
 - Saturated and Unsaturated solutions compares the ability of a solute to dissolve.

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Neil E. Schore



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