

## Identifying Reaction Types & Balancing Equations

<b>Step 1: Name the type of chemical reaction</b>	<b>Step 2: Balance the following chemical equations:</b>
Ex. Combination Reaction	<u>2</u> N <sub>2</sub> + <u>6</u> H <sub>2</sub> $\square$ <u>4</u> NH <sub>3</sub>
1.	<u>    </u> KClO <sub>3</sub> $\square$ <u>    </u> KCl + <u>    </u> O <sub>2</sub>
2. .	<u>    </u> NaCl + <u>    </u> F <sub>2</sub> $\square$ <u>    </u> NaF + <u>    </u> Cl <sub>2</sub>
3.	<u>    </u> H <sub>2</sub> + <u>    </u> O <sub>2</sub> $\square$ <u>    </u> H <sub>2</sub> O
4.	<u>    </u> AgNO <sub>3</sub> + <u>    </u> MgCl <sub>2</sub> $\square$ <u>    </u> AgCl + <u>    </u> Mg(NO <sub>3</sub> ) <sub>2</sub>
5.	<u>    </u> AlBr <sub>3</sub> + <u>    </u> K <sub>2</sub> SO <sub>4</sub> $\square$ <u>    </u> KBr + <u>    </u> Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>
6.	<u>    </u> CH <sub>4</sub> + <u>    </u> O <sub>2</sub> $\square$ <u>    </u> CO <sub>2</sub> + <u>    </u> H <sub>2</sub> O
7.	<u>    </u> C <sub>3</sub> H <sub>8</sub> + <u>    </u> O <sub>2</sub> $\square$ <u>    </u> CO <sub>2</sub> + <u>    </u> H <sub>2</sub> O
8.	<u>    </u> FeCl <sub>3</sub> + <u>    </u> NaOH $\square$ <u>    </u> Fe(OH) <sub>3</sub> + <u>    </u> NaCl
9.	<u>    </u> P + <u>    </u> O <sub>2</sub> $\square$ <u>    </u> P <sub>2</sub> O <sub>5</sub>
10.	<u>    </u> Na + <u>    </u> H <sub>2</sub> O $\square$ <u>    </u> NaOH + <u>    </u> H <sub>2</sub>
11.	<u>    </u> Ag <sub>2</sub> O $\square$ <u>    </u> Ag + <u>    </u> O <sub>2</sub>
12.	<u>    </u> CO <sub>2</sub> + <u>    </u> H <sub>2</sub> O $\square$ <u>    </u> C <sub>6</sub> H <sub>12</sub> O <sub>6</sub> + <u>    </u> O <sub>2</sub>
13.	<u>    </u> K + <u>    </u> MgBr <sub>2</sub> $\square$ <u>    </u> KBr + <u>    </u> Mg
14.	<u>    </u> HNO <sub>3</sub> (aq) + <u>    </u> Ba(OH) <sub>2</sub> (aq) $\square$ <u>    </u> Ba(NO <sub>3</sub> ) <sub>2</sub> (aq) + <u>    </u> H <sub>2</sub> O (l)
15.	<u>    </u> C <sub>3</sub> H <sub>8</sub> (g) + <u>    </u> O <sub>2</sub> (g) $\square$ <u>    </u> CO <sub>2</sub> (g) + <u>    </u> H <sub>2</sub> O(g)
16.	<u>    </u> Al(s) + <u>    </u> Fe <sub>2</sub> O <sub>3</sub> (aq) $\square$ <u>    </u> Al <sub>2</sub> O <sub>3</sub> (aq) + <u>    </u> Fe(s)
17.	<u>    </u> Al(s) + <u>    </u> O <sub>2</sub> (g) $\square$ <u>    </u> Al <sub>2</sub> O <sub>3</sub> (s)
18.	<u>    </u> P <sub>4</sub> + <u>    </u> O <sub>2</sub> $\square$ 2 P <sub>2</sub> O <sub>5</sub>
19.	<u>2</u> CaCO <sub>3</sub> $\square$ <u>    </u> CaO + <u>    </u> CO <sub>2</sub>
20.	<u>    </u> H <sub>2</sub> + <u>    </u> O <sub>2</sub> $\square$ 2H <sub>2</sub> O

# Identifying Reaction Types And Balancing Chemical Equations

**Liying Dong**



## **Identifying Reaction Types And Balancing Chemical Equations:**

## **Identifying Reaction Types And Balancing Chemical Equations** Book Review: Unveiling the Power of Words

In some sort of driven by information and connectivity, the power of words has be much more evident than ever. They have the capability to inspire, provoke, and ignite change. Such could be the essence of the book **Identifying Reaction Types And Balancing Chemical Equations**, a literary masterpiece that delves deep in to the significance of words and their affect our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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