

Ideal Gas Law

$$PV = nRT$$

↑ ↑ ↑ ↑
L mol K

atm

K

$$R = 0.08206$$

$$\frac{L \cdot atm}{mol \cdot K}$$

$$1 L = 1000 mL$$

$$T_K = T_C + 273$$

$$1 atm = 760 torr = 101.3 kPa$$

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