

LOTS OF  
CONVERSIONS!Name GrovePeriod     Date     /    /    **8 • Why Do Hot Air Balloons Float?****THE IDEAL GAS LAW**

$$PV = nRT \text{ where}$$

P = pressure in atmosphere

V = volume in liters

n = number of moles of gas

R = Universal Gas Constant = 0.0821 L·atm/mol·K

T = Kelvin temperature

See  
scratch  
paper

- How many moles of oxygen will occupy a volume of 2.50 liters at 1.20 atm and ~~25°C~~ <sup>298 K</sup>?  

$$PV = nRT \quad n = \frac{PV}{RT} = \frac{1.20 \text{ atm} \cdot 2.50 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 298 \text{ K}} = 0.947 \text{ mol}$$
- What volume will 2.00 moles of nitrogen occupy at ~~720 mmHg~~ <sup>0.947 atm</sup> and ~~20°C~~ <sup>293 K</sup>?  

$$V = \frac{nRT}{P} = \frac{2.00 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 293 \text{ K}}{0.947 \text{ atm}} = 500 \text{ L}$$
- What pressure will be exerted by ~~25.0 g~~ <sup>0.568 mol</sup> of  $\text{CO}_2$  at temperature of ~~25°C~~ <sup>298 K</sup> and a volume of ~~500 mL~~ <sup>0.500 L</sup>?  

$$P = \frac{nRT}{V} = \frac{0.568 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 298 \text{ K}}{0.500 \text{ L}} = 0.704 \text{ atm}$$
- At what temperature will ~~5.00 g~~ <sup>0.170 mol</sup> of  $\text{Cl}_2$  exert a pressure of ~~900 mmHg~~ <sup>1.18 atm</sup> at a volume of ~~250 mL~~ <sup>0.250 L</sup>?  

$$T = \frac{PV}{nR} = \frac{1.18 \text{ atm} \cdot 0.250 \text{ L}}{0.170 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}} = 210 \text{ K}$$
- What is the ~~density~~ <sup>9/L</sup> of  $\text{NH}_3$  at ~~800 mmHg~~ <sup>1.05 atm</sup> and ~~25°C~~ <sup>298 K</sup>?  

$$V = 1.00 \text{ L} \quad n = ?$$
- If the density of a gas is 1.2 g/L at ~~745 mmHg~~ <sup>0.980 atm</sup> and ~~20°C~~ <sup>293 K</sup>, what is its molar mass?  $\frac{g}{\text{mol}} = \frac{1.20 \text{ g}}{\text{mol}}$   

$$V = 1 \text{ L}$$
- How many moles of nitrogen gas will occupy a volume of ~~347 mL~~ <sup>0.347 L</sup> at ~~6680 mmHg~~ <sup>8.79 atm</sup> and ~~27°C~~ <sup>300 K</sup>?  

$$n = \frac{PV}{RT} = \frac{8.79 \text{ atm} \cdot 0.347 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 300 \text{ K}} = 0.125 \text{ mol}$$
- What volume will ~~454 grams~~ <sup>225 mol</sup> (1 lb) of hydrogen occupy at 1.05 atm and ~~25°C~~ <sup>298 K</sup>?  

$$V = \frac{nRT}{P} = \frac{225 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 298 \text{ K}}{1.05 \text{ atm}} = 502 \text{ L}$$
- Find the number of grams of  $\text{CO}_2$  that exert a pressure of ~~285 mmHg~~ <sup>0.375 atm</sup> at a volume of 32.5 L and a temperature of ~~32°C~~ <sup>305 K</sup>.  

$$n = ?$$
- An elemental gas has a mass of 10.3 g. If the volume is 58.4 L and the pressure is ~~758 mmHg~~ <sup>0.997 atm</sup> at a temperature of ~~25°C~~ <sup>275.5 K</sup>, what is the gas? Find molar mass

# Ideal Gas Law Chemistry If8766 Answers

**J Spring**



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