

Name _____ Period _____

Human Reproduction Webquest

Part I: The Benefits of Sexual Reproduction

- A. Go to the following webpage: http://www.pbs.org/wgbh/nova/miracle/program_adv.html
- B. Click on the first link: **passing on your DNA** and watch the movie.
 1. Why is sexual reproduction important?
 2. What is the process of making gametes called?

Part II: Spermatogenesis

- C. Go to the following webpage: http://wps.aw.com/bc_martin_l_eap_4/40/10469/2680298.cwr/content/index.html
- D. Click on the first link: **spermatogenesis** and watch the entire movie.
 3. Where does spermatogenesis occur?
 4. During puberty what hormone stimulates cells to begin meiosis?
 5. How many sperm cells are created from each meiotic division?
- E. Open another window and type in the following URL:
<http://highered.mcgraw-hill.com/olc/dl/120112/anim0043.swf>
- F. Watch the animation. Draw the process of spermatogenesis. Use the image and word bank below to help you!

Spermatogonia
Mitosis
Meiosis I
Meiosis II
Primary Spermatocyte
Secondary Spermatocytes
Spermatids
Differentiation
Sperm Cell



- G. Complete the following questions / matching section. Continue to watch the animation as needed so that you can do this.
 6. Are spermatogonia diploid or haploid cells?
 7. After the first division, mitosis, one cell remains a spermatogonia while the other continues on to form sperm. Why do you think it might be important to have one cell remain as a spermatogonia? (HINT: Men start making sperm at puberty and continue for their entire lives!)

Human Reproduction Webquest Answer Key

K Morrison



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Human Reproduction Casey Rand, 2009 An introduction to human reproduction that discusses the reproductive systems of males and females DNA fertilization in uterine development childbirth and other related topics *Topics in Human Reproduction* Linda M. Chagnon, Jill Wolhandler, 1975 Discusses the function and development of the male and female reproductive organs fertilization pregnancy and birth **Human Reproduction**, 1981 **Human Reproduction** J. I. D. Sadow, A. P. Gulamhusein, 1980 **Human Reproduction; Biology and Social Change** Harold D. Swanson, 1974 **Human Reproduction**, 19?? **Human Reproduction** J. I. D. Sadow, 1980 *Human Reproduction*, 1974 **Biology of Human Reproduction. Report of a WHO Scientific Group, Etc** Scientific Group on the Biology of Human Reproduction (WORLD HEALTH ORGANIZATION), 1964 **Problems of Human Reproduction** Paul Bowman Popenoe, 2012-12-01

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web jan 26 2016 the solubility product or ksp is a way of showing how effective a chemical is at dissolving into an aqueous solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then ionized Ca(OH)_2 when calcium is put in water reacts as follows $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$

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web k sp of calcium hydroxide 1 author j m mccormick last update october 13 2013 introduction an equilibrium constant K is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of K

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

web question experiment 8 determining ksp for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 hci molarity 0.050 m volume of filtered ca oh 2 solution ml 5.00 5.00 3.71 13.75 equivalence point volume ml oh in saturated ca oh 2 m ca in saturated ca oh 2 m ksp of ca oh 2 average ksp of ca oh 2

the chemical equation used was $\text{Ca(OH)}_2 \text{ (s)} + 2\text{HCl (aq)} \rightarrow \text{CaCl}_2 \text{ (aq)} + 2\text{H}_2\text{O (l)}$ the solubility product was determined to be $K_{sp} \text{ Ca(OH)}_2$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore

$$K_{\text{Ca}_2\text{PO}_4} = \frac{[\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2}{K_{\text{sp}}}$$

web lab report 8 determinin of calcium hydroxide the purpose of this was to titrat e a saturated Ca(OH)_2 soultion with a st
andard hcl solution and find the ksp or the oh concentration was f ound and then the concentration was found and u for this
experiment a titration of calcium hydroxide with hydrochlo

web question determining the ksp of calciumhydroxide Ca(OH)_2 s Ca^{2+} aq 2OH^{-} aq for this experiment the titration of calciumhydroxide with hydrochloric acid wasinvestigated determining the ksp of calcium hydroxide Ca(OH)_2 s Ca^{2+} aq 2OH^{-} aq for this experiment the titration of calcium

web question chem 182 experiment 8 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of Ca(OH)_2 is represented in equation form

soluble in water a saturated aqueous as shown below

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for ca was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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web plasma membrane homework symbiosis practice worksheet benzoin condensation grade a determining the ksp of calcium hydroxide after reading the lab complete items a b c and d title purpose chemicals and equipment and summary of procedure as described on page 10 of exp 1 on an 8 1 2 x 11 sheet of paper

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web we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the Ca^{2+} and OH^- concentrations in solution

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web calculate the ksp for calcium hydroxide $\text{Ca}^{2+} \text{OH}^-$ 2.0×10^{-2} ksp 5×10^{-4} find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is 5×10^{-6}

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