

[illegible]

2.50 mol
 H_2
1.00 mol
 N_2

Abstract

- ### ३. मरु-संस्कृत संज्ञा

☞ কোলাস পানিসূত্র লবণের দ্রাব্যতা আপনাত্মা কৃষ্ণির সাথে সাথে এক পর্যায়ে কৃষ্ণি না পেয়ে হ্রাস পায়। এর কারণ, আপনাত্মা কৃষ্ণির সাথে সাথে দ্রাব্যতা কৃষ্ণি পেতে থাকলেও একটি নির্দিষ্ট আপনাত্মার কোলাস পানি দ্রুত লবণ নিষ্কৃতিত হয়। তখন দ্রাব্যতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে দ্রবীভূত হওয়ার সময় বিয়োজন প্রক্রিয়াটি তাপস্বরী বিক্রিয়া কিন্তু 32.4°C আপনাত্মার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C আপনাত্মার উপরে আপনাত্মা কৃষ্ণিজনিত কারণে দ্রাব্যতা হ্রাস ঘটে।

w = SMV
 = $(0.5 \times 98 \times 0.25) \text{ g}$
 = 12.25 g
 অতএব, শুষ্কত H₂SO₄ এর ভর
 = 12.25 g

উদাহরণের A-পাত্রে অবস্থিত,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
 দ্রবণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
 দ্রবণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$

$$\text{[3]} \quad \text{HNO}_3 \rightarrow \text{H}_2\text{SO}_4 \text{ की मात्रा ज्ञात करें} = \left(0.5 \times \frac{250}{1000} \right) \text{ mol}$$

$$= 0.125 \text{ mol}$$

$\text{NaOH-H: Na}_2\text{CO}_3 \text{ এর মোল সংখ্যা} = \left(0.11 \times \frac{200}{1000} \right) \text{ mol}$
 $= 0.022 \text{ mol}$

अण्विहित समीकरण: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
अधिकतम शुद्ध भार, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

[illegible]

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 0.077 mol " " " " (0.077) mol " "
 " " " " " 0.077 mol "

অনশিট H_2SO_4 এর পরিমাণ = $(0.125 - 0.077) \text{ mol}$
 = 0.048 mol

পাত্র-A এর দ্রবণের সাথে পাত্র-B এর দ্রবণের বিক্রিয়ার ফলে ০.৫ mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লিক।

উদাহরণ ২ একজন ছাত্র ল্যাবরেটরিতে X-পদার্থিতে 10 mg এর ত্রুটি পরিমাণ এবং Y-পদার্থিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর দ্রবণ তৈরি করে।

References

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- উদ্ভীপকের এলিভটির মৌল্যত্ব ঘনমাত্রা নির্ণয় করো।
- উদ্ভীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

२. नए व्यक्तियों के लिए

উদাহরণ ১: নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেকট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

১৩। সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সম হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য ০.৫ এর কম হলে অপোলার হয়। যেহেতু Cl_1 এর তড়িৎ ঋণাত্মকতা ৩.০ ও C_2 অপূতে তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_1 বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লেখ্য মৌলিক অণুর ক্ষেত্রে এতদূর থাকবে।

গ) উদ্দীপকে বর্ণিত এসিডটি হলো $2\% \left(\frac{W}{V}\right) HCl$

2% $\left(\frac{W}{V}\right)$ HCl এর অর্থ হলো—

100 mL. મુશળ HCl એનું ઘન 2g

$$\therefore 1000 \text{ mL} = \left(\frac{2 \times 1000}{100} \right) \text{ g} = 20 \text{ g}$$

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Bowker

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Bowker' text.

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