

## DATA SHEET

|                                                               |                                                     |
|---------------------------------------------------------------|-----------------------------------------------------|
| Avogadro constant, $N_A$ .....                                | $6.022 \times 10^{23} \text{ mol}^{-1}$             |
| Volume of 1 mole ideal gas: at 100 kPa and                    |                                                     |
| at 0°C (273.15 K) .....                                       | 22.71 L                                             |
| at 25°C (298.15 K) .....                                      | 24.79 L                                             |
| Ionisation constant for water at 25°C (298.15 K), $K_w$ ..... | $1.0 \times 10^{-14}$                               |
| Specific heat capacity of water .....                         | $4.18 \times 10^3 \text{ J kg}^{-1} \text{ K}^{-1}$ |

## Some useful formulae

$$\text{pH} = -\log_{10}[\text{H}^+]$$

$$\Delta H = -m C \Delta T$$

## Some standard potentials

|                                                                      |             |                                                  |         |
|----------------------------------------------------------------------|-------------|--------------------------------------------------|---------|
| $\text{K}^+ + \text{e}^-$                                            | $\text{vs}$ | $\text{K(s)}$                                    | -2.94 V |
| $\text{Ba}^{2+} + 2\text{e}^-$                                       | $\text{vs}$ | $\text{Ba(s)}$                                   | -2.91 V |
| $\text{Ca}^{2+} + 2\text{e}^-$                                       | $\text{vs}$ | $\text{Ca(s)}$                                   | -2.87 V |
| $\text{Na}^+ + \text{e}^-$                                           | $\text{vs}$ | $\text{Na(s)}$                                   | -2.71 V |
| $\text{Mg}^{2+} + 2\text{e}^-$                                       | $\text{vs}$ | $\text{Mg(s)}$                                   | -2.36 V |
| $\text{Al}^{3+} + 3\text{e}^-$                                       | $\text{vs}$ | $\text{Al(s)}$                                   | -1.68 V |
| $\text{Mn}^{2+} + 2\text{e}^-$                                       | $\text{vs}$ | $\text{Mn(s)}$                                   | -1.18 V |
| $\text{H}_2\text{O} + \text{e}^-$                                    | $\text{vs}$ | $\frac{1}{2}\text{H}_2(\text{g}) + \text{OH}^-$  | -0.83 V |
| $\text{Zn}^{2+} + 2\text{e}^-$                                       | $\text{vs}$ | $\text{Zn(s)}$                                   | -0.76 V |
| $\text{Fe}^{2+} + 2\text{e}^-$                                       | $\text{vs}$ | $\text{Fe(s)}$                                   | -0.44 V |
| $\text{Ni}^{2+} + 2\text{e}^-$                                       | $\text{vs}$ | $\text{Ni(s)}$                                   | -0.24 V |
| $\text{Sn}^{2+} + 2\text{e}^-$                                       | $\text{vs}$ | $\text{Sn(s)}$                                   | -0.14 V |
| $\text{Pb}^{2+} + 2\text{e}^-$                                       | $\text{vs}$ | $\text{Pb(s)}$                                   | -0.13 V |
| $\text{H}^+ + \text{e}^-$                                            | $\text{vs}$ | $\frac{1}{2}\text{H}_2(\text{g})$                | 0.00 V  |
| $\text{SO}_4^{2-} + 4\text{H}^+ + 2\text{e}^-$                       | $\text{vs}$ | $\text{SO}_2(\text{aq}) + 2\text{H}_2\text{O}$   | 0.16 V  |
| $\text{Cu}^{2+} + 2\text{e}^-$                                       | $\text{vs}$ | $\text{Cu(s)}$                                   | 0.34 V  |
| $\frac{1}{2}\text{O}_2(\text{g}) + \text{H}_2\text{O} + 2\text{e}^-$ | $\text{vs}$ | $2\text{OH}^-$                                   | 0.40 V  |
| $\text{Cu}^+ + \text{e}^-$                                           | $\text{vs}$ | $\text{Cu(s)}$                                   | 0.52 V  |
| $\frac{1}{2}\text{I}_2(\text{s}) + \text{e}^-$                       | $\text{vs}$ | $\text{I}^-$                                     | 0.54 V  |
| $\frac{1}{2}\text{I}_2(\text{aq}) + \text{e}^-$                      | $\text{vs}$ | $\text{I}^-$                                     | 0.62 V  |
| $\text{Fe}^{3+} + \text{e}^-$                                        | $\text{vs}$ | $\text{Fe}^{2+}$                                 | 0.77 V  |
| $\text{Ag}^+ + \text{e}^-$                                           | $\text{vs}$ | $\text{Ag(s)}$                                   | 0.80 V  |
| $\frac{1}{2}\text{Br}_2(\text{l}) + \text{e}^-$                      | $\text{vs}$ | $\text{Br}^-$                                    | 1.08 V  |
| $\frac{1}{2}\text{Br}_2(\text{aq}) + \text{e}^-$                     | $\text{vs}$ | $\text{Br}^-$                                    | 1.10 V  |
| $\frac{1}{2}\text{O}_2(\text{g}) + 2\text{H}^+ + 2\text{e}^-$        | $\text{vs}$ | $\text{H}_2\text{O}$                             | 1.23 V  |
| $\frac{1}{2}\text{Cl}_2(\text{g}) + \text{e}^-$                      | $\text{vs}$ | $\text{Cl}^-$                                    | 1.36 V  |
| $\frac{1}{2}\text{Cr}_2\text{O}_7^{2-} + 7\text{H}^+ + 3\text{e}^-$  | $\text{vs}$ | $\text{Cr}^{3+} + \frac{7}{2}\text{H}_2\text{O}$ | 1.36 V  |
| $\frac{1}{2}\text{Cl}_2(\text{aq}) + \text{e}^-$                     | $\text{vs}$ | $\text{Cl}^-$                                    | 1.40 V  |
| $\text{MnO}_4^- + 8\text{H}^+ + 5\text{e}^-$                         | $\text{vs}$ | $\text{Mn}^{2+} + 4\text{H}_2\text{O}$           | 1.51 V  |
| $\frac{1}{2}\text{F}_2(\text{g}) + \text{e}^-$                       | $\text{vs}$ | $\text{F}^-$                                     | 2.89 V  |

# Hsc Chemistry Data Sheet

**JG Myers**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located on the right side of the slide, partially overlapping the light blue bar.

## **Hsc Chemistry Data Sheet:**

## Reviewing **Hsc Chemistry Data Sheet**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Hsc Chemistry Data Sheet**," an enthralling opus penned by a highly acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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