

Chapter 8 Chemical Reactions Chapter Review Answers

1. The evolution of energy, the production of a gas, the formation of a precipitate, and a color change.
2. A correctly written chemical equation must represent the known facts, contain the correct formulas for the reactants and products, and satisfy the law of conservation of mass.
3. a. It is a number that appears in front of a formula in a chemical equation.
b. The number of atoms of each type in the formula is multiplied by the coefficient.
4. methane + oxygen \rightarrow carbon dioxide + water;
 $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$;
 $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$
5. A chemical equation provides the number of molecules, moles, and atoms of each element or compound composing the reactants and products as well as the mole ratios of the reactants to each other and to the products.
6. Word and formula equations have only qualitative meaning; they do not express the relative quantities of a reaction's reactants or products.
7. a. a solution in water
b. a substance that accelerates a chemical reaction but can be recovered unchanged
c. a chemical reaction in which the products re-form the original reactants
8. a. KOH d. CCl_4
b. $\text{Ca}(\text{NO}_3)_2$ e. MgBr_2
c. Na_2CO_3
9. (1) Balance different types of atoms one at a time; (2) balance types of atoms that appear only once on each side of the equation first; (3) balance as single units any polyatomic ions that appear on both sides of the equation; and (4) balance H atoms and O atoms last.
10. a. 6N
b. 4H, 2O
c. 4H, 4N, 12O
d. 2Ca, 4O, 4H
e. 3Ba, 6Cl, 18O
f. 5Fe, 10N, 30O
g. 12Mg, 8P, 32O
h. 4N, 16H, 2S, 8O
i. 12Al, 18Se, 72O
j. 12C, 32H
11. a. $2\text{ZnS}(\text{s}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{ZnO}(\text{s}) + 2\text{SO}_2(\text{g})$
b. $2\text{HCl}(\text{aq}) + \text{Ba}(\text{OH})_2(\text{aq}) \rightarrow \text{BaCl}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
c. $2\text{HNO}_3(\text{aq}) + \text{Ca}(\text{OH})_2(\text{aq}) \rightarrow \text{Ca}(\text{NO}_3)_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
12. a. Solid zinc sulfide reacts with oxygen gas to form solid zinc oxide and sulfur dioxide gas.
b. When solid calcium hydride is added to water, aqueous calcium hydroxide and hydrogen gas are formed.
c. Aqueous silver nitrate mixed with aqueous potassium iodide produces a precipitate of silver iodide and aqueous potassium nitrate.
13. a. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
b. $2\text{Al} + \text{Fe}_2\text{O}_3 \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$
c. $\text{Pb}(\text{CH}_3\text{COO})_2 + \text{H}_2\text{S} \rightarrow \text{PbS} + 2\text{CH}_3\text{COOH}$
14. a. LiO_2 is an incorrect formula; $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$
b. H_2Cl_2 is an incorrect formula; $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
c. MgO_2 is an incorrect formula, and the equation as written is not balanced; $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
d. I is an incorrect formula for iodine, and the equation is not balanced; $2\text{NaI} + \text{Cl}_2 \rightarrow 2\text{NaCl} + \text{I}_2$
15. a. $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
b. $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$
c. $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
16. $\text{CH}_4(\text{g}) + 4\text{Cl}_2(\text{g}) \rightarrow \text{CCl}_4(\text{l}) + 4\text{HCl}(\text{g})$
17. a. $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
b. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
c. $2\text{Li} + \text{Cl}_2 \rightarrow 2\text{LiCl}$
d. $\text{Ca} + \text{I}_2 \rightarrow \text{CaI}_2$
18. (1) In a synthesis reaction, two or more substances combine to form a new substance: $\text{A} + \text{X} \rightarrow \text{AX}$.
(2) In a decomposition reaction, a single compound undergoes a reaction that produces two or more simpler substances: $\text{AX} \rightarrow \text{A} + \text{X}$.
(3) In a single-displacement reaction, one element replaces another element in a compound:
 $\text{A} + \text{BX} \rightarrow \text{AX} + \text{B}$ or $\text{Y} + \text{BX} \rightarrow \text{BY} + \text{X}$

Health Chemistry 8 Chapter Review Answers

Rod Powers



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Chemistry Karen Timberlake, 2003 *Chemistry An Introduction to General Organic and Biological Chemistry* now in its eighth edition makes chemistry exciting by showing why important concepts are relevant to the lives and future careers of readers. The new design, digital images, photos, Career Focus features, and macro to micro art enhance the new edition while it retains the many features that have made this book so successful. The writing as always is exceptionally friendly. Each section contains sample problems that develop readers' critical thinking skills. This edition also contains more conceptual problems than ever before and has been redesigned to accommodate new styles of learning and teaching with a wide variety of pedagogical tools. Health and Environmental Notes throughout the book highlight topics that are relevant to readers' lives and are ideal for classroom discussion. Explore Your World activities in each chapter make chemistry exciting, relevant, and non-threatening. *Introduction to General, Organic, and Biochemistry* Morris Hein, Scott Pattison, Susan Arena, Leo R.

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Environmental Engineering James R. Mihelcic, Julie B. Zimmerman, 2014-01-13 Environmental Engineering Fundamentals Sustainability Design presents civil engineers with an introduction to chemistry and biology through a mass and energy balance approach ABET required topics of emerging importance such as sustainable and global engineering are also covered Problems similar to those on the FE and PE exams are integrated at the end of each chapter Aligned with the National Academy of Engineering's focus on managing carbon and nitrogen the 2nd edition now includes a section on advanced technologies to more effectively reclaim nitrogen and phosphorous Additionally readers have immediate access to web modules which address a specific topic such as water and wastewater treatment These modules include media rich content such as animations audio video and interactive problem solving as well as links to explorations Civil engineers will gain a global perspective developing into innovative leaders in sustainable development

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