

A-Level Chemistry - Year 2

OCR Chemistry A - Paper 1 June - 2019

Question - 16

16 Sir Humphrey Davy discovered several elements including sodium, potassium, magnesium, calcium and strontium.

(a) Explain which block in the Periodic Table sodium and magnesium belong to.

(1)

(b) A sample of magnesium, $A_r = 24.305$, is found to consist of three isotopes. The accurate relative isotopic masses and % abundances of two of the isotopes are shown in the table.

Isotope	Relative isotopic mass	% abundance
^{24}Mg	23.985	78.99%
^{26}Mg	24.986	10.00%

Determine the relative isotopic mass of the third isotope of magnesium in the sample.

Give your answer to 5 significant figures.

relative isotopic mass = _____ (2)

OCR

International Baccalaureate

Tuesday 4 June 2019 – Afternoon

A Level Chemistry A

443201 Periodic table, elements and physical chemistry

Time allowed: 1 hour 15 minutes

Useful facts:

1. The data booklet is provided at the end of the paper.

Useful facts:

1. A periodic table is provided at the end of the paper.



Please write clearly in black ink. Do not write in the brackets.

Centre number

Candidate number

First name(s)

Last name

INSTRUCTIONS

- 1. Use black ink. You may use an HB pencil for graphs and diagrams.
- 2. Answer all the questions.
- 3. Where appropriate, your answers should be supported with writing. Marks may be given for a correct method even if the answer is incorrect.
- 4. Write your answer to each question in the space provided. If additional space is required, use the back page(s) at the end of this booklet. The question numbers must be clearly shown.

INFORMATION

- 1. The total mark for this paper is 100.
- 2. The marks for each question are shown in brackets ().
- 3. Quality of extended responses will be assessed in questions marked with an asterisk (*).
- 4. This document consists of 10 pages.

F321 Chemistry June 2014 Paper

Raffaella Di Napoli

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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