

Test Quantities in Chemical Reactions SCH3U #13

Name: _____

___K ___T ___C ___A Assessment of Learning (Product)
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Multiple Choice: Circle the most correct answer. (10K)

1. What is the molar mass of AuCl_3 ?

- A) 96 g B) 130 g C) 232.5 g D) 303.6 g

2. Determine the mass in grams of 0.0489 mol cobalt.

- A) 0.00812 g B) 0.294 g C) 2.88 g D) 0.000830 g

3. The number of molecules of hydrogen peroxide found in 3.15 mol H_2O_2 is:

- A) 3 B) 1.90×10^{23} C) 3.80×10^{23} D) 7.60×10^{23}

4. The percentage of oxygen, by mass, in $\text{Mg}_3(\text{PO}_4)_2$ is:

- A) 15.9% B) 23.6% C) 27.7%
 D) 48.7% E) None of the above

5. An unknown gas is found to consist of 10.48 g of nitrogen and 11.96 g of oxygen. What is the percentage composition of this gas?

- A) 28.6% N, 71.4 % O B) 46.7%N, 53.3% O
 C) 53.3% N, 46.7% O D) 74.5% N, 25.5% O

6. A compound contains 38.7% K, 13.9% N, and 47.4% O by mass. What is the empirical formula of the compound?

- A) KNO_3 B) $\text{K}_2\text{N}_2\text{O}_3$ C) KNO_2 D) K_2NO_3 E) K_4NO_3

7. If the molar mass of a hydrocarbon is 26.0 g/mol, and its empirical formula is CH , its molecular formula is

- A) CH B) C_2H_2 C) C_3H_3 D) C_4H_4 E) C_5H_5

8. Consider the following reaction: $\text{Ca}(\text{OH})_{2(s)} + \text{Al}(\text{SO}_4)_{3(s)} \rightarrow \text{CaSO}_{4(s)} + \text{Al}(\text{OH})_{3(s)}$

When balanced, the mole ratio for calcium hydroxide reacting with aluminum sulphate is

- A) 1:1 B) 2:1 C) 2:3 D) 1:2 E) 3:1

9. The limiting reactant in a chemical reaction is the _____ that is _____.

- A) Reactant, in excess B) Product, left over
 C) Reactant, used up D) Product, used up

10. Consider the following unbalanced reaction: $\text{Cs} + \text{H}_2\text{O} \rightarrow \text{H}_2 + \text{CsOH}$

Which statement about the above reaction is true?

- A) 2 g of Cs will react with 2 moles of water

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