



## Kawasaki KZ550 GT550 service manual #ch



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# Kawasaki Gt550 1994 Manual

**Kawasaki Jūkōgyō Kabushiki Kaisha.  
Motorcycle Division**



## **Kawasaki Gt550 1994 Manual:**

WALNECK'S CLASSIC CYCLE TRADER, JUNE 1994 Causey Enterprises, LLC,      *Suzuki Service Manual GT550 ,1973*

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fraction  $x$  is the ratio of moles of one substance in a mixture to the total number of moles of all substances for a mixture of two substances  $c_e a$  and  $c_e b$  the mole fractions of each would be written as

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web question 1 find the mole fraction of methanol  $\text{CH}_3\text{OH}$  and water in a solution prepared by dissolving  $4.5 \text{ g}$  of alcohol in  $40 \text{ g}$  of  $\text{H}_2\text{O}$  molar mass of  $\text{H}_2\text{O}$  is  $18 \text{ g mol}^{-1}$  and molar mass of  $\text{CH}_3\text{OH}$  is  $32 \text{ g mol}^{-1}$  solution moles of  $\text{CH}_3\text{OH}$   $4.5 / 32 = 0.14 \text{ mol}$  moles of  $\text{H}_2\text{O}$   $40 / 18 = 2.2 \text{ mol}$  therefore according to the equation

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web 1 in the given mixture the sum of all the mole fractions is equal to one  $x_a + x_b = 1$  the multiplication of the mole fraction by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction mole fraction is not dependent on the temperature

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web nov 6 2023  $x_a \text{ mol } a / \text{mol } a + \text{mol } b$  and  $x_b \text{ mol } b / \text{mol } a + \text{mol } b$  note in the given mixture the sum of all the mole fractions present is equal to one  $x_a + x_b = 1$  when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is a unitless and dimensionless expression image will be uploaded soon

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web solution total mass of the answer  $100 \text{ g}$  mass of benzene  $30 \text{ g}$  therefore mass of carbon tet  $100 - 30 = 70 \text{ g}$  benzene  $s$  molar

mass  $\text{C}_6\text{H}_6$  6 12 6 1 g mol 1 78 g mol 1 hence  $\text{C}_6\text{H}_6$  s number of moles  $\frac{30.78 \text{ mol}}{0.3846 \text{ mol}}$  molar mass of carbon tet  $\text{CCl}_4$  12 4 355 154 g mol 1

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web solution moles of  $\text{CH}_3\text{OH}$  5 5 32 0 17 mole moles of  $\text{H}_2\text{O}$  40 18 2 2 moles therefore according to the equation mole fraction of  $\text{CH}_3\text{OH}$  0 17 2 2 0 17 mole fraction of  $\text{CH}_3\text{OH}$  0 073 to solve more examples on mole fraction formula and practice more question please visit byju s com

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal kbr that contains 2 92 kg of water 1 molality moles solute divided by kilograms solute 0 449 mol kg x 2 92 kg x 1 31108 mol of kbr 2 moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g kbr

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web calculate the mass percent mole fraction molality and molarity of hexane  $\text{C}_6\text{H}_{14}$  d 0 660 g  $\text{cm}^3$  when it is mixed with heptane  $\text{C}_7\text{H}_{12}$  d 0 684 g  $\text{cm}^3$  in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web since the mole fraction of the o 2 is 0 050 it contributes 0 050 or if you will 5 0 of the total pressure thus total pressure is 0 21 0 050 4 2 atm problem 9 a sample of 1 43 g of helium and an unweighed quantity of o 2 are mixed in a flask at room temperature

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