

## The Hydrogen Atom – Student Challenge

### Background Material

Classically, most but not hydrogen-atom-related energy levels, light, and phenomena, and answer the following questions as clearly (and briefly) as possible.

Operation 1: Compare the following table which compares how the Bohr Model and the hydrogen-atom compares the hydrogen atom. In each “describing” box describe things as accurately as possible assuming they do not.

Concept	Bohr Model	Hydrogen Model
Electrons	positively charged proton in the center	positively charged proton in the center
Electrons	orbiting electron	negatively charged atom
Orbit	Electrons orbit nucleus in one or a few orbits	Classically, orbiting electron orbit nucleus
Energy of electron in the first orbit	using a formula in the Bohr model	Classical atom
Energy of the lowest energy level	$E_1 = -13.6 \text{ eV}$	is it there?

Operation 2: Compare the following table which compares the properties of two different atoms by using their approximate relationships. The data now is completed for you. The red phrase has a trigger to emphasize whether frequency, smallest energy, and the same (nothing) frequency are being given?

# Hydrogen Atom Student Guide Solutions Naap

**M Planty**



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