

# Holt Modern Chemistry Review

## CHAPTERS 16: REACTION ENERGY

The following pages contain the bulk (but not all) of the information for the chapter 16 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and **NOTHING** else, you should at least be able to **PASS** the test. <sup>\*\*\*\*\*</sup> Test items will be recall, examples, and/or application of this content. <sup>\*\*\*\*\*</sup>

### OUTCOMES

- Collaborate with peer(s) to understand chemistry content (C.C)
- Communicate chemistry content to teacher and peer(s) (C.C)
- 16.1: Distinguish between heat and temperature and their units (T.A.R)
- 16.1: Calculate heat problems using common specific heats on wkbk, pg. 505 & txbk, pg. 533 (E.A.P.R)
- 16.1: Explain enthalpy change & endothermic / exothermic reactions (graphs on wkbk, pg. 506-509 & txbk, pg. 536-537) (T.A.R)

### 16.1: THERMOCHEMISTRY

- **Vocabulary**
  - **thermochemistry** — the branch of chemistry that is the study of the energy changes that accompany chemical reactions and changes of state
  - **calorimeter** — a device used to measure the energy as heat absorbed or released in a chemical or physical change
  - **temperature** — a measure of how hot (or cold) something is; specifically, a measure of the average kinetic energy of the particles in an object; the unit should be in kelvins
  - **joule** — the unit used to express energy; equivalent to the amount of work done by a force of 1 N acting through a distance of 1 m in the direction of the force; abbreviation, J
  - **heat** — energy is always transferred from higher-temperature objects to lower-temperature objects until thermal equilibrium is reached; the energy transferred between objects that are at different temperatures; the SI unit for this as well as all other forms of energy is joule, J
  - **specific heat** — the quantity of heat required to raise a unit mass of homogeneous material 1 K or 1°C in a specified way given constant pressure and volume
  - **enthalpy change** — the amount of energy released or absorbed as heat by a system during a process at constant pressure
  - **enthalpy of reaction** — the amount of energy released or absorbed as heat during a chemical reaction
  - **thermochemical equation** — an equation that includes the quantity of energy as heat released or absorbed during the reaction as written
- **Chapter Highlights**
  - Thermochemistry is the study of the changes in energy that accompany chemical reactions and physical changes.
  - A thermochemical equation is an equation that includes the quantity of energy released or absorbed as heat during the reaction as written.
  - A battery (battery acid) in a digital pH meter is an example of chemical energy being used in a chemistry lab.
  - An enthalpy change is the amount of energy absorbed as heat by a system in a process carried out at constant pressure.
  - The enthalpy of reaction is the enthalpy change that occurs during a chemical reaction.
  - The enthalpy change is negative for exothermic reactions and positive for endothermic reactions.
  - Compounds with highly negative enthalpies of formation tend to be stable; compounds with highly positive or only slightly negative enthalpies of formation tend to be unstable.
  - The standard molar enthalpy of formation is the enthalpy change that occurs when one mole of a compound is formed from its elements in their standard states at 25°C and 1 atm.
  - The enthalpy change that occurs in a combustion reaction is called the enthalpy of combustion.
  - Enthalpies of reaction can be calculated by using enthalpies of formation of reactants and products.

# Holt Modern Chemistry Reaction Energy Review Answers

**G Orfield**



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